

A new local density functional for main-group thermochemistry, transition metal bonding, thermochemical kinetics, and noncovalent interactions

Yan Zhao^{a)} and Donald G. Truhlar

*Department of Chemistry, University of Minnesota, Minneapolis, Minnesota 55455-0431
and Supercomputing Institute, University of Minnesota, Minneapolis, Minnesota 55455-0431*

(Received 31 July 2006; accepted 28 September 2006; published online 15 November 2006)

We present a new local density functional, called M06-L, for main-group and transition element thermochemistry, thermochemical kinetics, and noncovalent interactions. The functional is designed to capture the main dependence of the exchange-correlation energy on local spin density, spin density gradient, and spin kinetic energy density, and it is parametrized to satisfy the uniform-electron-gas limit and to have good performance for both main-group chemistry and transition metal chemistry. The M06-L functional and 14 other functionals have been comparatively assessed against 22 energetic databases. Among the tested functionals, which include the popular B3LYP, BLYP, and BP86 functionals as well as our previous M05 functional, the M06-L functional gives the best overall performance for a combination of main-group thermochemistry, thermochemical kinetics, and organometallic, inorganometallic, biological, and noncovalent interactions. It also does very well for predicting geometries and vibrational frequencies. Because of the computational advantages of local functionals, the present functional should be very useful for many applications in chemistry, especially for simulations on moderate-sized and large systems and when long time scales must be addressed. © 2006 American Institute of Physics.

[DOI: 10.1063/1.2370993]

I. INTRODUCTION

In the early days of density functional theory,¹ all density functionals were local. Local functionals may depend on the local spin density,^{2,3} its gradient⁴⁻⁹ or Laplacian,¹⁰ or even spin kinetic energy density approximated in terms of the kinetic energy of Kohn-Sham spin-orbitals.¹¹⁻¹³ Unfortunately functionals depending on the gradient of the spin density were sometimes called nonlocal, but that semantic error appears to be disappearing. More recently, nonlocal density functionals have been widely employed, and in many cases they greatly improve the accuracy.¹⁴⁻¹⁷ The present article returns to the formulation of local density functionals.

It is important to clarify the significance of local density functionals. Some researchers have voiced the opinion that nonlocal density functionals are to be tolerated as an interim “fix” until the true density functional is derived/discovered/developed. This is a misunderstanding. The theorem¹⁸ that exact density functionals exist does not apply to local density functionals; thus we must allow a density functional to be nonlocal if it is to be exact.^{19,20} Nonlocal density functionals include hybrid functionals that depend on Hartree-Fock exchange.¹⁴ Thus local functionals are important not because they somehow represent a more theoretically justified solution, rather they are important for practical reasons because calculations on large complex systems may employ specialized algorithms (including density fitting, also called resolution of the identity) that are tens or hundreds of times faster if one employs local density functionals than if one employs nonlocal ones.²¹⁻³⁰ For example, we found that a single-point

energy calculation for a C₁₀₄H₃₀N₄ fullerene-porphyrin complex³¹ with a local functional employing density fitting is 15 times faster (17 vs 250 h) than a hybrid density functional calculation on the same system [using GAUSSIAN03 (Ref. 32)]. Furthermore, calculations are impractical for many solids when Hartree-Fock exchange is included,³³ unless some special techniques³⁴⁻³⁷ are employed.

A second reason to study local density functionals is their usefulness for modeling the bonding of metallic elements. Two recent systematic studies of highly unsaturated systems containing metal atoms, one for metal-metal bonds³⁸ and one for metal-ligand bonds,³⁹ have shown that most functionals with more than about 10% Hartree-Fock exchange fail badly for a large number of bonds involving transition metals, which is understandable in that the density-based exchange functionals do a better job than Hartree-Fock exchange in accounting for static correlation.^{9,40-43} Even more recently we have presented a new functional,⁴⁴ called M05, that performs well for such cases even with 28% Hartree-Fock exchange and also performs well for main-group chemistry, barrier heights, and noncovalent interactions. Nevertheless, we are still interested in local density functionals because of their cost advantages discussed in the previous paragraph.

In the present article, we present a local functional that has better general performance than the most popular hybrid functional, B3LYP. The new functional is called M06-L. We assess this new functional by applying it, along with 11 previous local functionals [BP86,^{5,7} BLYP,^{7,8} BB95,^{7,11} G96LYP,^{8,45} PBE,⁴⁶ mPWPW,^{16,47} VSXC,¹² HCTH,⁴⁸ OLYP,⁴¹ τ -HCTH,⁴⁹ and TPSS (Ref. 13)] and three previous

^{a)}Electronic mail: yzhao@chem.umn.edu

nonlocal functionals [B3LYP,^{7,14,50} TPSSh,⁵¹ and M05 (Ref. 44)], to the data in 22 diverse databases. In fact, the average performance of the M06-L functional over the 22 energetic databases considered here is better than that of any local or hybrid functional that we know.

Most of the databases used in the present work have been introduced and described in previous work. One new database introduced here is a database of five metal-atom excitation energies containing two main-group neutral metals, two neutral transition metals, and one transition metal cation. The addition of this data set is prompted by the increasing attention being paid by many workers to the relative energies of spin states of transition metal systems because of the importance of spin states for structures, properties, and chemical reactivities of organometallic complexes and for functional nanotechnology.⁵²⁻⁶² Holthausen made a systematic study⁶³ of the ability of density functionals to predict $3d-4s$ excitation energies in transition metal cations and found that some functionals, even though performing well for main-group atomization energies, show large errors for transition metal atomic excitation energies. Because we seek a functional that is accurate for both main-group and transition metal chemistries, our small database has both types of atoms. We note that our transition metal excitation energy comparisons take account of scalar relativistic effects,⁶⁴ whereas those of Holthausen do not, even though such effects are sometimes large.

The paper is organized as follows: Sec. II presents our databases. Section III gives computational details. Section IV discusses the theory and parametrization of the new functionals. Section V presents results and discussion, including test for energies, geometries, and frequencies not used in training, and Sec. VI concludes the paper.

II. ENERGETIC DATABASES

Our general notation for databases is XN/V , where X is an acronym for the type of data, N is the number of data, and V is a version number (sometimes omitted if there has only been one version). The reason for version numbers is that historically we have sometimes corrected (if possible) or eliminated data whose reliability has been disproved or credibly challenged. All data in Subsections II.A. through II.H. are pure electronic energies, i.e., zero-point energies and thermal vibrational-rotational energies have been removed by methods discussed previously,^{17,65-68} but nuclear repulsion is included. Since the databases are based on experimental or accurate data, the values in our database correspond to relativistic values. Thus our calculations must include relativistic corrections as well, where these are not negligible; this is discussed in Sec. III.B.

II.A. MGAE109/05 test set

The MGAE109/05 test set¹⁷ consists of 109 atomization energies (AEs) for main-group compounds. We always give the mean errors in atomization energies on a per bond basis because that makes comparison between different test sets more portable. In the past, as some workers increased the size of their test sets, they tended to add larger molecules,

and the resulting increase in mean errors due to increasing the average number of bonds could not be distinguished from the increase in mean errors due to the added diversity of the test molecules. To make it possible for readers to compare other workers' values to our mean errors on a per bond basis, we always compute the mean errors in atomization energies by computing the mean error per molecule and then dividing by the average number of bonds per molecule in the test set. The latter value is 4.71 for MGAE109/05.

II.B. Ionization potential, electron affinity, and proton affinity test sets

The zero-point-exclusive ionization potential (IP) and electron affinity (EA) test sets are called IP13/3 and EA13/3, respectively, and they have been explained and employed in our previous papers.^{17,44,65,68} PA8 is a database of eight zero-point-exclusive proton affinities.⁶⁹ These three databases involve only small molecules and atoms.

II.C. Barrier height databases

The HTBH38/04 database contains 38 transition state barrier heights for 19 hydrogen transfer (HT) reactions, 18 of which involve radicals as reactant and product. They are taken from previous papers.^{17,67,70} NHTBH38/04 consists of three databases containing 38 transition state barrier heights for nonhydrogen-transfer (NHT) reactions, and these databases have been presented in a previous paper.⁷⁰ The individual databases contain 12 barrier heights for heavy-atom transfer reactions, 16 barrier heights for nucleophilic substitution (NS) reactions, and 10 barrier heights for non-NS unimolecular and association reactions.

II.D. Noncovalent interaction databases

Recently we developed several databases, in particular, HB6/04,⁷¹ CT7/04,⁷¹ DI6/04,⁷¹ WI7/05,⁶⁸ and PPS5/05,⁶⁸ for various kinds of noncovalent interactions. HB6/04 is a hydrogen bond database that consists of the equilibrium binding energies of six hydrogen bonding dimers. The CT7/04 database consists of binding energies of seven charge transfer complexes. The DI6/04 database contains the binding energies of six dipole interaction complexes. The WI7/05 database consists of the binding energies of seven weak interaction complexes, all of which are bound by dispersion interactions. The PPS5/05 database consists of binding energies of five $\pi-\pi$ stacking complexes. We also test the density functionals against a benchmark database of 22 noncovalent interactions of biological importance; it is called S22, and it was developed by Jurecka *et al.*⁷² The S22 database includes both small and large systems, with the largest system containing 30 atoms.

II.E. Transition metal dimer and transition metal-ligand databases

We employ three databases involving molecules containing transition metals. One³⁸ is for atomization energies of transition metal-transition metal dimers, and it is called the TMAE9/05 database; it contains the bond energies of Ag_2 ,

Cr₂, CuAg, Cu₂, Mo₂, Ni₂, V₂, ZrV, and Zr₂. The other,³⁹ called MLBE21/05, is for the metal-ligand bond energies in organometallic and inorganometallic complexes, and it contains 21 metal-ligand bond energies. We also test the density functionals against a benchmark database of 18 reaction energies involving 3*d* transition metals;⁷³ we called it 3dTMRE18/06, and it was developed by Furche and Perdew.⁷⁴

II.F. Alkyl bond dissociation energy (ABDE) database

The next database contains four *R*–*X* bond dissociation energies D_e (*R*=methyl and isopropyl and *X*=CH₃ and OCH₃) and is called the ABDE4/05 database. The reference D_0 values are taken from a recent paper by Izgorodina *et al.*,⁷⁵ and we used the B3LYP/6-31G(*d*) zero-point vibrational energies scaled with a scale factor of 0.9806 (Ref. 76) to obtain the accurate D_e values.

II.G. π -system databases

We employ three databases for π systems, namely, π IE3/06, PA-CP5/06, and PA-SB5/06. The π IE3/06 database contains three isomeric energy differences between allene and propyne as well as higher homologs (which correspond to cumulenes and polyenes).^{69,77} PA-CP5/06 is a database of the proton affinities of five conjugated polyenes. PA-SB5/06 is a database of the proton affinities of the five conjugated Schiff bases. All structures for these molecules are given in the supporting information.⁷⁸

II.H. Excitation energy database

MAEE5 is a database of five metal-atom excitation energies containing two main-group neutral metals (Be and Mg), two neutral transition metals (Mn and Pd), and one transition metal cation (Cu⁺). One reason that this database is small is to avoid the complication of mixed symmetry states that makes density functional theory (DFT) predictions ambiguous for many open-shell transition metal atoms.^{63,73,74}

II.I. Bond length and frequency databases

MGBL19 is a database of 19 experimental bond lengths of 15 small main-group molecules, and the experimental bond lengths are taken from a previous compilation by Hamprecht *et al.*⁴⁸ MLBL13/05 is a previously³⁹ developed database of bond lengths of 13 metal-ligand compounds. F36/06 is a new database of 36 harmonic frequencies of the 13 molecules in a previous vibrational zero-point energy database (ZPE13/99).^{79,80} The experimental harmonic frequencies are taken from a previous compilation by Martin *et al.*⁸¹ with one exception; since there is no reliable experimental harmonic frequency for the umbrella mode of the NH₃ molecule, the best estimate for this mode is taken from a CCSD(T)/cc-pVQZ calculation by Martin *et al.*^{81,82}

II.J. Supporting information

The 3dTMRE18/06 database is completely described in Ref. 74, and the ABDE4/06, MAEE5/06, MGBL19/06, and

F36/06 databases will be presented below. The other 19 databases are given in Tables S1–S9 of the supporting information.⁷⁸

III. COMPUTATIONAL METHODS

III.A. Geometries and basis sets

All calculations for the AE6, MGAE109/05, IP13/3, EA13/3, HTBH38/04, and NHTBH38/05 databases are single-point calculations at QCISD/MG3 geometries, where QCISD is quadratic configuration interaction with single and double excitations,⁸³ and MG3 is the modified^{84,85} G3Large (Ref. 86) basis set. The MG3 basis set,⁸⁴ also called G3LargeMP2,⁸⁵ is the same as 6-311++G(3*d*2*f*,2*d**f*,2*p*) (Refs. 87 and 88) for H–Si, but improved⁸⁶ for P–Ar. Geometries for the PA8/06 database are at the MP2(full)/6-31G(2*d**f*,*p*) level of theory.

Geometries for all molecules in the HB6/04, CT7/04, DI6/04, and WI7/05 noncovalent databases and the (C₂H₄)₂ and (C₂H₂)₂ dimers in the PPS5/05 database are optimized at the MC-QCISD/3 level, where MC-QCISD is the multicoefficient QCISD method.^{66,89} The geometries for the benzene dimers in the PPS5/05 database are taken from Sinnokrot and Sherrill.⁹⁰ The methods used to obtain geometries⁷² for the S22 database are specified in the supporting information, and all DFT calculations for the S22 database employ the 6-31+G(*d*,*p*) basis set.

Geometries for all molecules in the ABDE4/05 database are optimized at the B3LYP/6-31G(*d*) level, and they are given in the supporting information of a previous paper.⁷⁵ The 6-311+G(3*d**f*,2*p*) basis set is used for the calculations of ABDEs; this choice is made for the purpose of comparison with the previous results.

Geometries for the molecules in the transition metal (TMAE9/05) and metal-ligand (MLBE21/05) bond energy databases are optimized consistently with each level of theory. We used the triple-zeta-quality (TZQ) basis set³⁸ for the calculations on the molecules in these two databases. For the 3dTMRE18 database, we employed a quadruple-zeta-quality basis set, QZVP, developed by Weigend *et al.*⁹¹ and TPSS/QZVP geometries.⁷⁴

To test the functionals for the MAEE5 database, we employed the aug-cc-pVQZ basis set for Be,⁹² Mg,⁹² and Cu⁺.⁹³ The basis set for Mn is the QZVP basis of Weigend *et al.*⁹¹ and that for Pd is the AVQZ basis set from Quintal *et al.*⁹⁴

The MG3S basis set was employed for all the calculations on the MGBL19 and F36 databases.

III.B. Relativistic effects

In all cases, experimental data in the database are zero-point exclusive, but (like all experimental data) they include relativistic effects. Since our DFT calculations are nonrelativistic, we must add relativistic effects to our computed Born-Oppenheimer energies in order to compare with experiment.

We divide relativistic effects into two types: scalar and vector. Scalar relativistic effects are generally very small for H-Cl, and these small effects are not included for main-group compounds in the present study. Scalar relativistic effects for

systems in the 3dTMRE18/06 database were treated in exactly the same way as in Ref. 74. We also include scalar relativistic effects for the excitation energies of Mn and Cu⁺ (MAEE5/06), and for the 4*d* transition metals in the TMAE9/05, MLBE21/05, and MAEE5/06 databases. In most cases we take account of the scalar relativistic effect simply by using a relativistic effective core potential. For the 4*d* metals, we use the relativistic effective core potential of Stevens *et al.*⁹⁵ to replace the 1*s*²2*s*²2*p*⁶3*s*²3*p*⁶3*d*¹⁰ core, and the 4*s*, 4*p*, 5*s*, and 4*d* electrons are treated explicitly. For calculations on Pd atom, we employed the Stuttgart-Dresden relativistic effective core potential⁹⁶ (MWB28) to replace the 28 core electrons, and the 18 valence electrons are treated explicitly. The scalar relativistic contributions for the excitation energy of the Mn and Cu⁺ are taken from previous work.^{64,97}

The vector relativistic effect is included by adding spin-orbit coupling. In all of the calculations presented in this paper, the spin-orbit stabilization energy was added to atoms and open-shell molecules for which it is nonzero, as described previously.^{38,39,84} Note that spin-orbit coupling vanishes for closed-shell species, atoms in *S* states, molecules in Σ states, and singlet and doublet molecules in *A* or *B* states.

III.C. Counterpoise correction

For noncovalent complexes, we perform calculations with and without the counterpoise corrections^{98,99} for basis set superposition error (BSSE).

III.D. Software

All calculations were performed with a locally modified version of the GAUSSIAN03 program.³²

IV. THEORY AND PARAMETRIZATION

In a recent review,³³ Scuseria and Staroverov summarized the six strategies that have been widely employed for designing density functionals: (1) local density approximation (LDA), (2) density-gradient expansion (DGE), (3) constraint satisfaction, (4) modeling the exchange-correlation hole, (5) empirical fits, and (6) mixing Hartree-Fock and approximate DFT exchange. Our M05 (Ref. 44) and M05-2X (Ref. 17) functionals were constructed with strategies (3), (5), and (6), and several of the key ideas were drawn from earlier work by Becke, Perdew, and others.^{3,6,11,46,49,100,102,103} The M05 functional form is well suited to employing a high percentage of Hartree-Fock exchange. When we tried optimizing a local functional we found that the M05 functional form is not flexible enough to provide a local functional with broad accuracy. From the literature^{12,51,104} and from one of our previous assessments,¹⁰⁵ one can see that the VSXC local functional¹² performs very well for thermochemistry (and that will be reconfirmed in the present study). However, further studies^{106–108} also show that VSXC gives very discouraging performance for noncovalent interactions, and the barrier heights, although better than those from other local functionals, are not as good as one would desire. In contrast the M05-2X functional has excellent performance for noncovalent interactions and barrier heights, and M05 has reason-

able performance for noncovalent interactions and barriers. These observations motivated us to combine the functional forms of VSXC and M05, and this combination provides the functional form adopted for M06-L. Interestingly, the VSXC functional (also called VS98) was developed using a density matrix expansion to model the exchange-correlation hole, a strategy of type (4). For the M06-L functional, we also enforce the uniform- electron-gas limit; therefore the M06-L functional is constructed based on strategies (3)–(5). We note that the original parametrization of VSXC did not satisfy the uniform-electron-gas (UEG) limit. The authors allowed this limit to be violated (by ~5% for exchange, 67% for same-spin correlation, and 30% for opposite-spin correlation) to improve performance for atomic and molecular systems. Here we enforce this limit because our functional form is sufficiently flexible that enforcing it does not degrade performance for molecular systems.

The key issue in optimizing functionals is whether or not one has a functional form that captures the physically significant dependences. Experience, both ours and that of others,^{12,109} has shown that simply increasing the number of parameters does not necessarily yield a qualitatively more accurate functional or one that fits a more diverse set of data; the functional form is the key element.

Some preliminary studies showed that some terms in the VSXC functional cause an integration grid problem pointed by Johnson *et al.*,¹⁰⁶ so we set the coefficients for these terms to zero. Before we describe the details of the new functional, we define some basic variables and functions. The spin density is called ρ_σ , and the following equations define the reduced spin density gradient x_σ , spin kinetic energy density τ_σ , a working variable z_σ , and two working functions γ and h :

$$x_\sigma = \frac{|\nabla\rho_\sigma|}{\rho_\sigma^{4/3}}, \quad \sigma = \alpha, \beta, \quad (1)$$

$$\tau_\sigma = \frac{1}{2} \sum_i^{\text{occup}} |\nabla\Psi_{i\sigma}|^2, \quad (2)$$

$$z_\sigma = \frac{\tau_\sigma}{\rho_\sigma^{5/3}} - C_F, \quad C_F = \frac{3}{5}(6\pi^2)^{2/3}, \quad (3)$$

$$\gamma(x_\sigma, z_\sigma) = 1 + \alpha(x_\sigma^2 + z_\sigma), \quad (4)$$

$$h(x_\sigma, z_\sigma) = \left(\frac{d_0}{\gamma(x_\sigma, z_\sigma)} + \frac{d_1 x_\sigma^2 + d_2 z_\sigma}{\gamma_\sigma^2(x_\sigma, z_\sigma)} + \frac{d_3 x_\sigma^4 + d_4 x_\sigma^2 z_\sigma + d_5 z_\sigma^2}{\gamma_\sigma^3(x_\sigma, z_\sigma)} \right). \quad (5)$$

Note that σ denotes the component of electron spin angular momentum and takes on two values: $\alpha(+1/2)$ and $\beta(-1/2)$.

It will be noticed in Secs. IV.A and IV.B that the new density functional depends on both τ_α and τ_β , not just on their sum. We use the standard language by which “spin” as an adjective means “spin resolved.” Thus we say that the functional depends on the spin kinetic energy densities rather

than saying that it depends on the kinetic energy density, just as the functional depends on spin densities and their gradients, not just their sums.

The M06-L functional may be classified as a meta-generalized gradient-approximation (meta-GGA), where meta denotes the dependence on spin kinetic energy density.

IV.A. Meta-GGA exchange functional

The exchange functional used in the M05 functional⁴⁴ was an extension of earlier functionals developed by Becke,^{6,11,101,102} and Perdew *et al.*¹⁵ The M05 functional form is well suited to employing a high percentage of Hartree-Fock exchange. In order to make a better local functional for the present work we added additional terms based on the relatively successful local exchange functional of Voorhis and Scuseria,^{12,110} which is based in part on earlier work by Negele and Vautherin.¹¹¹ In particular, the M06-L exchange functional is given by

$$E_X^{\text{M06-L}} = \sum_{\sigma} \int dr [F_{X\sigma}^{\text{PBE}}(\rho_{\sigma}, \nabla \rho_{\sigma}) f(w_{\sigma}) + \varepsilon_{X\sigma}^{\text{LSDA}} h_X(x_{\sigma}, z_{\sigma})], \quad (6)$$

where $h_X(x_{\sigma}, z_{\sigma})$ is defined in Eq. (5) [with $d_5=0$ in $h_X(x_{\sigma}, z_{\sigma})$], $F_{X\sigma}^{\text{PBE}}(\rho_{\sigma}, \nabla \rho_{\sigma})$ is the exchange energy density of the Perdew-Burte-Ernzerhof⁴⁶ (PBE) exchange model, $\varepsilon_{X\sigma}^{\text{LSDA}}$ is the local spin density approximation (LSDA) for exchange

$$\varepsilon_{X\sigma}^{\text{LSDA}} = \frac{3}{2} \left(\frac{3}{4\pi} \right)^{1/3} \rho_{\sigma}^{4/3}, \quad (7)$$

and $f(w_{\sigma})$ is the spin kinetic energy density enhancement factor

$$f(w_{\sigma}) = \sum_{i=0}^m a_i w_{\sigma}^i, \quad (8)$$

where the variable w_{σ} is a function of t_{σ} , and t_{σ} is a function of the spin kinetic energy density τ_{σ} and spin density ρ_{σ} ,

$$w_{\sigma} = (t_{\sigma} - 1)/(t_{\sigma} + 1), \quad (9)$$

$$t_{\sigma} = \tau_{\sigma}^{\text{LSDA}} / \tau_{\sigma}, \quad (10)$$

where

$$\tau_{\sigma}^{\text{LSDA}} \equiv \frac{3}{10} (6\pi^2)^{2/3} \rho_{\sigma}^{5/3}. \quad (11)$$

IV.B. Meta-GGA correlation functional

The M06-L correlation functional is based on the M05 correlation functional, again augmented by terms from Voorhis and Scuseria. The M05 correlation functional builds on work of Perdew and Wang,³ Becke,¹¹ Perdew,⁴⁷ and Stoll *et al.*¹⁰⁰ In the correlation functional, we treat the opposite-spin and parallel-spin correlation differently.

The opposite-spin correlation energy of our new functional is expressed as

$$E_C^{\alpha\beta} = \int e_{\alpha\beta}^{\text{UEG}} [g_{\alpha\beta}(x_{\alpha}, x_{\beta}) + h_{\alpha\beta}(x_{\alpha\beta}, z_{\alpha\beta})] dr, \quad (12)$$

where $g_{\alpha\beta}(x_{\alpha}, x_{\beta})$ is defined as

$$g_{\alpha\beta}(x_{\alpha}, x_{\beta}) = \sum_{i=0}^n c_{C\alpha\beta,i} \left(\frac{\gamma_{C\alpha\beta}(x_{\alpha}^2 + x_{\beta}^2)}{1 + \gamma_{C\alpha\beta}(x_{\alpha}^2 + x_{\beta}^2)} \right)^i \quad (13)$$

and $h_{\alpha\beta}(x_{\alpha\beta}, z_{\alpha\beta})$ is defined in Eq. (5), with $x_{\alpha\beta}^2 \equiv x_{\alpha}^2 + x_{\beta}^2$ and $z_{\alpha\beta} \equiv z_{\alpha} + z_{\beta}$.

For parallel spins,

$$E_C^{\sigma\sigma} = \int e_{\sigma\sigma}^{\text{UEG}} [g_{\sigma\sigma}(x_{\sigma}) + h_{\sigma\sigma}(x_{\sigma}, z_{\sigma})] D_{\sigma} dr, \quad (14)$$

where $g_{\sigma\sigma}(x_{\sigma})$ is defined as

$$g_{\sigma\sigma}(x_{\sigma}) = \sum_{i=0}^n c_{C\sigma\sigma,i} \left(\frac{\gamma_{C\sigma\sigma} x_{\sigma}^2}{1 + \gamma_{C\sigma\sigma} x_{\sigma}^2} \right)^i \quad (15)$$

and $h_{\sigma\sigma}(x_{\sigma}, z_{\sigma})$ is defined in Eq. (5). D_{σ} is the self-interaction correction factor

$$D_{\sigma} = 1 - \frac{x_{\sigma}^2}{4(z_{\sigma} + C_F)}. \quad (16)$$

Significantly, D_{σ} vanishes for any one-electron system.

Note that $e_{\alpha\beta}^{\text{UEG}}$ and $e_{\sigma\sigma}^{\text{UEG}}$ in Eqs. (12) and (14) are the UEG correlation energy density for the antiparallel-spin and parallel spin cases, and they can be extracted from the total UEG correlation energy density.¹⁰⁰ The total correlation energy of the new correlation functional is given by

$$E_C = E_C^{\alpha\beta} + E_C^{\alpha\alpha} + E_C^{\beta\beta}. \quad (17)$$

The values of the two nonlinear parameters in Eqs. (13) and (15) are taken from previous work,^{17,44}

$$\gamma_{C\alpha\beta} = 0.0031, \quad \gamma_{C\sigma\sigma} = 0.06. \quad (18)$$

The values of the three nonlinear parameters α_x , $\alpha_{C\alpha\beta}$, and $\alpha_{C\sigma\sigma}$ in Eq. (5), as employed in Eqs. (6), (12), and (14), are taken from a different previous work.¹²

IV.C. Optimization of the new local meta-GGA

All parameter optimizations were carried out in a self-consistent fashion. The parameters a_i in Eq. (6), $c_{C\alpha\beta,i}$ in Eq. (12), and $c_{C\sigma\sigma,i}$ in Eq. (14), along with the parameters d_i in $h_X(x_{\sigma}, z_{\sigma})$ [Eq. (6)], $d_{C\alpha\beta,i}$ in $h_{\alpha\beta}(x_{\alpha\beta}, z_{\alpha\beta})$ [Eq. (12)] and $d_{C\sigma\sigma,i}$ in $h_{\sigma\sigma}(x_{\sigma}, z_{\sigma})$ [Eq. (14)], were determined by fitting to the data in the training set. To obtain the correct UEG limit, we enforce the following constraints

$$a_0 + d_0 = 1, \quad (19)$$

$$c_{C\alpha\beta,i} + d_{C\alpha\beta,i} = 1, \quad (20)$$

$$c_{C\sigma\sigma,i} + d_{C\sigma\sigma,i} = 1. \quad (21)$$

We optimized the parameters in M06-L against accurate data to approximately minimize (see below for details) a training function F defined by

TABLE I. Optimized parameters in the M06-L Methods.

Parameters	M06-L					
	a_i	$c_{C\alpha\beta,i}$	$c_{C\sigma\sigma,i}$	d_i	$d_{C\alpha\beta,i}$	$d_{C\sigma\sigma,i}$
0	3.987 756E-01	6.042 374E-01	5.349 466E-01	6.012 244E-01 ^a	3.957 626E-01 ^a	4.650 534E-01
1	2.548 219E-01	1.776 783E+02	5.396 620E-01	4.748 822E-03	-5.614 546E-01	1.617 589E-01
2	3.923 994E-01	-2.513 252E+02	-3.161 217E+01	-8.635 108E-03	1.403 963E-02	1.833 657E-01
3	-2.103 655E+00	7.635 173E+01	5.149 592E+01	-9.308 062E-06	9.831 442E-04	4.692 100E-04
4	-6.302 147E+00	-1.255 699E+01	-2.919 613E+01	4.482 811E-05	-3.577 176E-03	-4.990 573E-03
5	1.097 615E+01					
6	3.097 273E+01					
7	-2.318 489E+01					
8	-5.673 480E+01					
9	2.160 364E+01					
10	3.421 814E+01					
11	-9.049 762E+00					

^aDetermined by a constraint, as explained in Sec. IV.C.

$$\begin{aligned}
 F = & \text{RMSEPB(MGAE109)} + \text{RMSE(IP13)} + \text{RMSE(EA13)} \\
 & + \text{RMSE(PA8)} + \text{RMSE(BH76)} \\
 & + 10 \times \text{RMSE(NCCE31)} + \text{RMSE(TMML30)} \\
 & + \text{RMSE(ABDE4)} + \text{RMSE(AE17)} + \text{RMSE}(\pi 13), \quad (22)
 \end{aligned}$$

where the terms have the following meaning: RMSEPB is the root-mean-squared error (RMSE) per bond for the MGAE109 database. The second term is the RMSE for the IP13 database. The third term is the RMSE for the EA13 database. RMSE(PA8) is the RMSE for the PA8 database of eight proton affinities. RMSE(BH76) is the RMSE for the 76 barrier heights in the HTBH38 and NHTBH38 databases. RMSE(NCCE31) is the RMSE for the 31 noncovalent complexation energies in the HB6/04, CT7/04, DI6/04, WI7/05,

and PPS5/05 databases. RMSE(ABDE4) is the RMSE for the bond dissociation energies in the ABDE4/05 database. RMSE(TMML30) is the RMSE in the bond dissociation energies in the TMAE9/05 and MLBE21/05 databases. RMSE(AE17) is the RMSE for the 17 total atomic energies¹¹² of the atoms from H to Cl. RMSE(π 13) is the RMSE of the 13 energetic data in the π IE3/06, PA-CP5/06, and PA-SB5/06 databases.

We minimize the training function with respect to these parameters in a self-consistent way by solving the Fock-Kohn-Sham equation using the basis sets and geometries described in Sec. III.A but subject to two constraints: $|c_{C\alpha\beta,0}| < 5$ and $|c_{C\sigma\sigma,0}| < 5$. Because of these constraints, the training function is only approximately minimized, but the functional is more physical than for an unconstrained optimization.

TABLE II. Tested density functionals.

Method	Year	Ref(s).	Exchange				Correlation			
			$\rho_\sigma, \nabla\rho_\sigma$	X	τ	UEG	$\rho_\sigma, \nabla\rho_\sigma$	τ_σ	SCorF	UEG
Local										
BP86	1988	5 and 7	B88	0	No	Yes	P86	No	No	Yes
BLYP	1988	7 and 8	B88	0	No	Yes	LYP	No	Yes	No
BB95	1996	7 and 11	B88	0	No	Yes	B95	Yes	Yes	Yes
G96LYP	1996	8 and 45	G96	0	No	Yes	LYP	No	Yes	No
PBE	1996	46	PBE	0	No	Yes	PBE	No	No	Yes
mPWPW ^a	1998	47 and 16	mPW	0	No	Yes	PW91	No	No	Yes
VSXC ^b	1998	12	VSXC	0	Yes	No	VSXC	Yes	Yes	No
HCTH	1998	48	HCTH	0	No	No	HCTH	No	No	No
OLYP	2001	41	OPTX	0	No	No	LYP	No	Yes	No
τ -HCTH	2002	49	τ -HCTH	0	Yes	No	τ -HCTH	No	No	No
TPSS	2003	13	TPSS	0	Yes	Yes	TPSS	Yes	Yes	Yes
M06-L	2006	17	M06-L	0	Yes	Yes	M06-L	Yes	Yes	Yes
Nonlocal										
B3LYP	1994	7, 14, and 50	B88	20	No	No	LYP	No	Yes	No
TPSSh	2003	51	TPSS	10	Yes	Yes	TPSS	Yes	Yes	Yes
M05	2005	44	M05	28	Yes	Yes	M05	Yes	Yes	Yes

^aAlso called mPWPW91.

^bAlso called VS98.

TABLE III. Mean errors^a [kcal/mol for ionization potentials (IPs), electron affinities (EAs), proton affinities (PAs), and kcal/mol per bond for atomization energies (AEs)].

Method	AE109		IP13		EA13		PA8		TMUE ^c
	MSEPB	MUEPB ^a	MSE ^b	MUE	MSE ^b	MUE	MSE ^b	MUE	
Local									
VSXC	-0.18	0.57	2.31	3.29	0.22	2.80	1.83	1.98	1.10
OLYP	-0.05	0.86	-1.50	2.66	3.23	3.57	2.30	2.30	1.35
M06-L	0.05	0.85	0.76	3.09	2.96	3.84	2.01	2.06	1.39
TPSS	0.63	1.03	1.80	3.11	0.51	2.31	2.67	2.67	1.43
τ -HCTH	-0.08	0.85	3.66	4.32	-0.88	2.61	3.17	3.17	1.46
HCTH	-0.31	1.05	4.85	5.46	-2.37	3.75	2.69	2.69	1.79
BLYP	-0.47	1.49	-0.41	4.87	-0.11	2.63	-0.69	1.53	1.90
mPWPW	1.72	2.01	2.93	4.15	-1.56	2.26	0.88	1.49	2.20
G96LYP	-1.39	1.96	-1.12	4.64	1.33	2.93	0.81	1.40	2.26
BB95	2.18	2.34	-0.55	3.34	0.21	1.99	0.23	1.65	2.36
PBE	2.80	3.03	2.11	3.58	-1.20	2.22	0.04	1.35	2.91
BP86	3.55	3.65	4.48	5.38	-4.23	4.23	0.35	1.37	3.73
Nonlocal									
M05	-0.01	0.53	-0.41	2.87	2.81	2.96	1.20	2.16	1.06
B3LYP	-0.69	0.91	3.58	4.72	-1.51	2.29	0.18	1.02	1.39
TPSSh	-0.12	0.98	1.96	3.17	1.40	2.81	2.78	2.78	1.45
Average ^d	0.51	1.48	1.63	3.91	0.05	2.88	1.36	1.98	1.85

^aMUEPB denotes mean unsigned error (MUE) per bond.

^bMSE denotes mean signed error.

^cTMUE denotes total MUE and it is defined as $TMUE = [MUEPB \times 109 + MUE(IP) \times 13 + MUE(AE) \times 13 + MUE(PA) \times 8] / 143$

^dIn all tables where the last row is "Average," it is the average of that column for all functionals in the table.

tion. The M05 and M05-2X functionals were each parametrized against 35 data and have been tested on 354 data. Examination of Eq. (22) shows that 314 data are used during the parametrization of M06-L.

All optimized parameters for M06-L are listed in Table I. Notice that $a_0 = 0.398\ 775\ 6$ and $d_0 = 0.601\ 224\ 4$. One can interpret this as meaning that the new exchange functional has 40% τ -enhanced PBE exchange (similar to M05 exchange) and 60% modified VSXC exchange. One can also obtain small errors with a larger fraction of modified VSXC exchange, but we saw no reason to increase this fraction arbitrarily. We note that the VSXC functional was originally parametrized without enforcing the uniform-electron-gas limit in order to improve empirical performance, but the present functional does not require this compromise of an exact constraint.

V. RESULTS AND DISCUSSION

We compare the results obtained by the new M06-L functional to those for 11 other local functionals and three hybrid functionals. Note that in the present paper, we are focusing on the performance of the local functionals, and readers can find the performance of many other hybrid functionals in our previous work^{17,68,80,113} and the work of others.^{114–116} We include three hybrid functionals in the present work simply to provide perspective. We specifically omit high Hartree-Fock exchange functionals such as PWB6K,⁶⁸ M05-2X,¹⁷ and BMK (Ref. 116) because they are well known to be inaccurate for many problems involving transition metals. Table II lists all 15 density functionals con-

sidered in this work. In each case we specify the year it was first published, the functional forms used for dependence on $\nabla\rho_\sigma$, whether or not the functional includes τ_σ in the exchange and correlation functional, and whether the correlation functional is self-correlation-free (SCorF). Table II also contains two columns (one for the exchange functional and one for the correlation functional) that tell whether or not the functional reduces to the correct uniform-electron-gas limit when $\nabla\rho_\sigma \rightarrow 0$ and $\tau_\sigma \rightarrow \tau_\sigma^{LSDA}$.

In most of the comparisons we will gauge the quality of the results by mean unsigned errors (MUEs), which are the averages of the absolute deviations of calculated values from database reference values, and by mean signed errors (MSEs), which are used to detect systematic deviations. However, for atomization energies we use MUE per bond (MUEPB) and MSE per bond (MSEPB) because, as discussed in Sec. II, this allows^{105,113,117} more transferable comparison between databases with different average sizes of molecules. To make the trends clearer, in every table that ends with a mean error column we will list the methods in increasing order of the values in the key (overall) error column, which is always the last column of a given table.

Section II presented 22 energetic databases (one in Sec. II.A, three in Sec. II.B, one in Sec. II.C, three in Sec. II.D, six in Sec. II.E, three in Sec. II.F, one in Sec. II.G, three in Sec. II.H, and one in Sec. II.I). Here we discuss them in the same order, but for discussion purposes we group them slightly differently (four in Sec. V.A, four in Sec. V.B, six in

TABLE IV. Mean errors for thermochemical kinetics. (The QCISD/MG3 geometries and MG3S basis set are used for calculations in this table.)

Methods	Heavy atom transfer ^a		Nuc. sub. ^b		Unimol. ^c		Hydrogen transfer ^d		AMUE ^e	MMUE ^h
	MSE ^e	MUE ^f	MSE ^e	MUE ^f	MSE ^e	MUE ^f	MSE ^e	MUE ^f		
					Local					
M06-L	-5.58	5.93	-3.58	3.58	0.04	1.86	-4.14	4.16	3.02	3.88
HCTH	-8.84	8.84	-2.71	2.71	-0.75	2.20	-5.41	5.47	3.36	4.81
VSXC	-7.44	7.44	-5.30	5.30	-0.91	2.40	-4.86	4.87	3.45	5.00
OLYP	-11.23	11.23	-2.73	2.73	-1.92	2.53	-5.59	5.63	3.92	5.53
τ -HCTH	-9.21	9.21	-5.71	5.71	-1.04	2.82	-6.82	6.87	4.04	6.15
G96LYP	-13.03	13.03	-5.80	5.80	-2.86	3.04	-6.25	6.26	4.75	7.03
BB95	-13.88	13.88	-6.36	6.36	-3.22	3.40	-8.14	8.14	5.29	7.94
MPWPW	-14.10	14.10	-7.45	7.45	-2.67	3.10	-8.43	8.43	5.38	8.27
BLYP	-14.66	14.66	-8.40	8.40	-3.38	3.51	-7.52	7.52	5.67	8.52
TPSS	-14.65	14.65	-7.75	7.75	-3.84	4.04	-7.71	7.71	5.72	8.54
PBE	-14.93	14.93	-6.97	6.97	-2.94	3.35	-9.32	9.32	5.65	8.64
BP86	-15.51	15.51	-6.91	6.91	-3.41	3.87	-9.16	9.16	6.05	8.86
					Nonlocal					
M05	-2.84	3.79	0.00	0.80	0.69	2.24	-1.20	1.93	2.06	2.19
B3LYP	-8.49	8.49	-3.25	3.25	-1.42	2.02	-4.13	4.23	3.10	4.50
TPSSh	-11.51	11.51	-5.78	5.78	-2.94	3.23	-5.97	5.97	4.57	6.62
Average	-11.06	11.15	-5.25	5.30	-2.04	2.91	-6.31	6.38	4.40	6.43

^aReference 12.^b"Nuc. sub." denotes nucleophilic substitution reactions (Ref. 16).^cThis denotes unimolecular and association reactions (Ref. 10).^dReference 38.^eMSE denotes mean signed error (kcal/mol).^fMUE denotes mean unsigned error (kcal/mol).^gAMUE is defined as $AMUE = [MUE(\Delta E, 38) + MMUE]/2$, where $MUE(\Delta E, 38)$ is the mean unsigned error for the energy of reactions for the 38 reactions involved in this table. AMUE is one measure of the quality of a method for kinetics.^hMMUE in this table is calculated by averaging the numbers in columns 3, 5, 7, and 9.

Sec. V.C, three in Sec. V.D, one in Sec. V.E, three in Sec. V.F, and one in Sec. V.G), after which we provide some tests for bond lengths and vibrational frequencies.

V.A. Thermochemistry: AE, IP, EA, and PA results

Table III summarizes the errors in AE, IPs, EAs, and proton affinities (PAs) for all tested functionals. Table III shows that the M05, VSXC, and M06-L methods give the best results for AE calculations.

OLYP, M06-L, and M06-L have the best performance for IP calculations, whereas BB95, mPWPW, and B3LYP give the best performance for EA calculations. For the calculation of proton affinities, B3LYP, PBE, and BP86 give the best performance

To compare their performance for thermochemistry, we defined the total MUE (TMUE) as the mean signed error over all 143 data in this table as follows:

$$TMUE = [MUEPB(AE) \times 109 + MUE(IP) \times 13 + MUE(EA) \times 13 + MUE(PA8)]/143. \quad (23)$$

If we use TMUE as a criterion of practical usefulness for main-group thermochemistry, Table III shows that M05 is the best functional, followed by VSXC and M06-L. The results in Table III confirm the point in Sec. IV that VSXC is the best local functional for main-group thermochemistry. It is encouraging that it is followed by the OLYP and M06-L

functionals; all three match or outperform both the most popular hybrid functional, B3LYP, and its local version, BLYP, although they do not outperform M05. We next turn our attention to 20 other databases to test for a local functional with good performance across the board.

V.B. Thermochemical kinetics

Table IV gives the mean errors for the HTBH38/04 and NHTBH38/04 databases with the MG3S basis set. We also tabulated a value of mean MUE (called MMUE) that is defined as 1/4 times the MUE for heavy-atom transfer barrier heights plus 1/4 times the MUE for S_N2 barrier heights plus 1/4 times the MUE for unimolecular and association barrier heights plus 1/4 times the MUE for hydrogen transfer barrier heights.

Table IV shows that the M05, M06-L, and VSXC methods give the best results for heavy-atom-transfer barrier height calculations. M05, OLYP, and HCTH have the best performance for nucleophilic substitution barrier height calculations. M06-L, B3LYP, and HCTH give the best performance for unimolecular and association barrier height calculations. The M05, M06-L, and B3LYP methods give the best performance for hydrogen transfer barrier height calculations, and they also give the lowest values of MMUE, which means they give the best overall performance for barrier height calculations.

Another quantity, average MUE or AMUE, is defined as

TABLE V. Mean errors for noncovalent databases (kcal/mol). (The MG3S basis set is used for calculations in this table.)

Method	HB6/04			CT7/04			DI6/04			WI7/05			PPS5/05			
	MUE ^a			MUE ^a			MUE ^a			MUE ^a			MUE ^a			
	No cp ^d	cp ^e	MMUE ^b	No cp ^d	cp ^e	MMUE ^b	No cp ^d	cp ^e	MMUE ^b	No cp ^d	cp ^e	MMUE ^b	No cp ^d	cp ^e	MMUE ^b	MMUE ^c
	Local															
M06-L	0.21	0.51	0.36	1.80	1.41	1.61	0.32	0.32	0.32	0.19	0.13	0.16	0.17	0.42	0.29	0.55
PBE	0.45	0.32	0.39	2.95	2.63	2.79	0.46	0.40	0.43	0.13	0.15	0.14	1.86	2.09	1.97	1.14
τ -HCTH	0.67	1.12	0.90	1.77	1.47	1.62	0.50	0.53	0.52	0.22	0.23	0.22	2.46	2.72	2.59	1.17
TPSS	0.45	0.82	0.63	2.20	1.86	2.03	0.52	0.56	0.54	0.19	0.26	0.22	2.53	2.78	2.66	1.22
mPWPW	0.57	0.96	0.77	2.25	1.89	2.07	0.56	0.59	0.57	0.24	0.32	0.28	2.69	2.96	2.83	1.30
HCTH	1.68	2.12	1.90	1.31	1.22	1.26	0.55	0.71	0.63	0.28	0.21	0.25	3.06	3.35	3.21	1.45
BP86	0.72	1.10	0.91	2.03	1.73	1.88	0.66	0.76	0.71	0.65	0.74	0.70	3.22	3.45	3.33	1.51
BLYP	1.18	1.56	1.37	1.67	1.42	1.54	1.00	1.18	1.09	0.45	0.53	0.49	3.58	3.79	3.69	1.63
BB95	1.83	2.21	2.02	1.48	1.27	1.38	1.18	1.35	1.27	0.57	0.66	0.62	2.96	3.18	3.07	1.67
VSXC	0.45	0.79	0.62	2.84	2.53	2.68	1.10	1.02	1.06	0.94	0.90	0.92	6.75	6.58	6.66	2.39
OLYP	3.60	4.09	3.84	1.57	1.58	1.57	2.35	2.53	2.44	0.38	0.44	0.41	4.72	4.96	4.84	2.62
G96LYP	2.95	3.30	3.13	1.20	1.28	1.24	2.57	2.74	2.65	1.37	1.47	1.42	5.19	5.41	5.30	2.75
	Nonlocal															
M05	0.58	0.53	0.55	0.68	0.30	0.49	0.23	0.24	0.23	0.14	0.06	0.10	1.12	1.34	1.23	0.52
TPSSh	0.41	0.80	0.60	1.44	1.16	1.30	0.49	0.58	0.54	0.18	0.26	0.22	2.46	2.72	2.59	1.05
B3LYP	0.60	0.93	0.76	0.71	0.54	0.63	0.78	0.94	0.86	0.31	0.39	0.35	2.95	3.17	3.06	1.13
Average	1.09	1.41	1.25	1.73	1.49	1.61	0.88	0.96	0.92	0.42	0.45	0.43	3.00	3.20	3.10	1.46

^aMUE denotes mean unsigned error.

^bMMUE=[MUE(cp)+MUE(no-cp)]/2.

^cMMMUE=[MMUE(HB)+MMUE(CT)+MMUE(DI)+MMUE(WI)+MMUE(PPS)]/5; HB: hydrogen bonding; CT: charge transfer; DI: dipole interaction; WI: weak interaction; PPS: π - π stacking.

^d“No cp” denotes the calculation without the counterpoise correction for the BSSE.

^e“cp” denotes the calculation with the counterpoise correction for the BSSE.

$$\text{AMUE} = [\text{MUE}(\Delta E, 38) + \text{MMUE}(\text{BH76})]/2, \quad (24)$$

where MUE($\Delta E, 38$) is the mean unsigned error in the energy of reactions for the 38 reactions in the HTBH38 and NHTBH38 databases. If one prefers to use AMUE as a criterion to justify the performance of a DFT method for thermochemical kinetics, the conclusions one draws are essentially the same as above; in particular, the rank order of the eleven best functionals in Table IV is the same using AMUE as using MMUE.

V.C. Noncovalent interactions

In general, noncovalent interactions are key elements determining many of the properties of soft materials.^{118–125} The mean errors for noncovalent interaction are listed in Tables V and VI. In Table V, we use “No cp” to denote calculations without the counterpoise correction for the BSSE, and we use “cp” to denote calculations that do include the counterpoise correction for the BSSE. In Tables V and VI, we also defined a mean MUE as follows:

$$\text{MMUE} = [\text{MUE}(\text{no cp}) + \text{MUE}(\text{cp})]/2. \quad (25)$$

This is a reasonable error criterion because the cp correction is sometimes an overestimate of BSSE, and because in practical work some calculations are carried out with cp corrections and some without. This is actually just a technical point in the present paper because the table shows that (because

our basis set is large enough) the conclusions are the same if we base them on cp, no cp, or MMUE results.

Table V shows that PBE, M06-L, and M05 give the best performance for calculating the binding energies of the hydrogen bonding dimers in the HB6/04 database. These three methods also give the best performance for calculating the binding energies of the dipole interaction complexes in the DI6/04 database.

Charge transfer complexes have long been recognized as a difficulty for local functionals;^{126,127} Table V confirms that most local functionals give very poor performance for calculating the binding energies for the complexes in the CT7/04 databases. M06-L, while not as accurate for this property as the hybrid M05 and B3LYP functionals, is the sixth best performing local functional in Table V, and its MUE of 1.61 kcal/mol is the same as the average (1.61 kcal/mol as given in the last row) over all 15 functionals in the table.

We note that weak interactions and π - π stacking interactions play a dominant role in stabilizing various biopolymers, for example, the double helix structure of DNA,^{128,129} and such interactions are also important for protein folding¹³⁰ and supramolecular design.^{131,132} Table V shows that the quality of M06-L for describing π - π stacking interactions is much better than all other local functionals and is also better than M05, B3LYP, and TPSSh. The best functional for noncovalent interactions is M05-2X,¹⁷ which is not included in

TABLE VI. Mean errors for the S22 noncovalent database of biological importance.

Method	MMUE-HB ^a	MMUE-D ^b	MMUE-mix ^c	MMMUE ^d
		Local		
M06-L	0.82	0.76	0.72	0.77
PBE	1.13	4.53	1.66	2.44
TPSS	1.59	5.62	2.36	3.19
mPWPW	1.86	5.95	2.55	3.45
τ -HCTH	1.76	6.75	2.67	3.73
BP86	2.02	6.41	2.94	3.79
BB95	3.54	5.43	2.97	3.98
HCTH	3.29	6.57	2.48	4.11
BLYP	2.94	7.43	3.45	4.60
G96LYP	5.11	10.55	5.50	7.05
OLYP	5.98	10.43	4.94	7.12
VSXC	1.27	15.12	5.18	7.19
		Nonlocal		
M05	1.26	3.16	1.09	1.83
TPSSh	1.41	5.42	2.22	3.01
B3LYP	1.77	6.22	2.64	3.54
Average	2.38	6.69	2.89	3.99

^aSeven hydrogen bonded complexes.

^bEight complexes dominated by dispersion-like interactions, including π - π stacking.

^cSeven mixed complexes, e.g., benzene \cdots H₂O.

^dAverage of these previous columns.

the present study because it is not designed to be a universal functional for treating transition metals as well as main-group chemistry.

The overall performance for noncovalent interactions can be judged by the mean MMUE, which is defined as

$$\text{MMMUE} = [\text{MMUE}(\text{HB}) + \text{MMUE}(\text{CT}) + \text{MMUE}(\text{DI}) + \text{MMUE}(\text{WI}) + \text{MMUE}(\text{PPS})]/5. \quad (26)$$

Notice that the five components in Eq. (26) place different requirements on a density functional. For example, high accuracy for charge transfer complexes is not well correlated with high accuracy for weak interactions. If we use MMMUE as a criterion to evaluate the overall performance of DFT methods for noncovalent interactions, we can see from Table V that M05, M06-L, and TPSSh are the best functionals in the present study.

We have recently tested 40 functionals against the S22 database, and we found that the six most accurate functionals were all hybrid functionals.¹⁰⁸ Therefore this is a difficult test for local functionals. Table VI presents results for the S22 database. Note that S22 is a large, diverse database that is not in the training set of the M06-L functional. It is encouraging that M06-L performs well for this database of biologically important interactions. In fact, it is more than a factor of 2 better than any other local functional, and it is also considerably more accurate than the three hybrid functionals in Table VI. The S22 database was not used in the training step.

V.D. Transition element bond energies

Metal-metal and metal-ligand bondings are very important in many application areas.^{52,60,133–146} Table VII summa-

TABLE VII. MUE (kcal/mol) for the TMAE9/05, MLBE21/05, and 3dTMRE18/06 databases. (The TZQ basis set is employed for the TMAE9/05 and MLBE21/05 databases. The QZVP basis set is employed for the 3dTMRE18/06 basis set.)

Method	TMAE9/05		MLBE21/05		3dTMRE18/06		MMUE
	MSE	MUE	MSE	MUE	MSE	MUE	
			Local				
M06-L	0.2	4.9	4.7	5.4	3.8	6.9	5.7
G96LYP	0.2	4.8	7.0	7.7	4.4	10.3	7.6
OLYP	-2.8	7.7	4.9	6.3	2.9	9.0	7.7
TPSS	-1.3	6.1	7.4	7.9	7.9	10.2	8.1
BLYP	4.8	5.3	9.0	9.6	5.5	10.6	8.5
VSXC	5.5	10.2	6.3	6.6	8.2	8.9	8.6
mPWPW	0.5	6.4	10.2	10.6	9.1	10.3	9.1
HCTH	11.3	11.9	7.3	7.6	4.5	8.0	9.2
BP86	5.6	7.6	11.8	12.2	8.9	10.3	10.0
PBE	3.9	7.7	11.7	12.1	10.3	10.8	10.2
BB95	9.5	9.5	12.6	12.9	12.4	12.5	11.7
τ -HCTH	20.2	20.2	9.0	9.2	11.6	11.9	13.8
			Nonlocal				
M05	-3.0	6.9	-0.7	5.5	-3.0	7.8	6.8
TPSSh	-11.0	11.0	2.1	5.5	1.0	9.7	8.7
B3LYP	-16.7	16.7	-0.6	6.0	-6.6	12.0	11.6
Average	1.8	9.1	6.9	8.3	5.4	10.0	9.1

rizes the results for the TMAE9/05, MLBE21/05, and 3dTMRE18/06 databases. For the TMAE9/05 database of bond energies of transition metal dimers, G96LYP, BLYP, and M06-L give the best results. For the MLBE21/05 database of metal-ligand compounds, TPSSh, M05, and M06-L give the best performance, and M06-L, M05, and HCTH perform best for the 3dTMRE18/06 database. Thus M06-L is the only functional in the top three for all these databases.

In Table VII, MMUE is the average of the MUE for the TMAE9/05, MLBE21/05, and 3dTMRE18/06 databases, and M06-L, M05, and G96LYP give the smallest MMUE. The good performance of M06-L on all three databases is encouraging, especially since 3dTMAE18/06 was not used during parametrization.

V.E. Trends in alkyl bond dissociation energies

Table VIII summarizes the results for the trends in R - X bond dissociation energies (BDEs) (R =Me and i -Pr; X =CH₃ and OCH₃). Table VII shows that BB95, PBE, and M05 give the best performance for this BDE database. The ABDE database was originally created in response to previous criticism^{75,147} of DFT that it performs more poorly for larger alkyl groups than for methyl. It is encouraging that the accuracy of M06-L degrades less rapidly than other functionals as the alkyl group size is increased from methyl to isopropyl.

V.F. Tests for π systems

Table IX summarizes the results for the energetic data in the π IE3/06, PA-CP5/06, and PA-SB5/06 databases. M05, PBE, and B3LYP give the best performance. M06-L per-

TABLE VIII. Alkyl bond dissociation energies (D_e , kcal/mol). [The B3LYP/6-31G(d) geometries are used in all calculations in this table. All DFT calculations in this table use the 6-311+G(3df,2p) basis set.]

Method	R-CH3		R-OCH3		MSE	MUE
	R=Me	R=i-Pr	R=Me	R=i-Pr		
Expt.	97.39	95.00	89.79	91.51		
	Local					
BB95	98.35	90.15	87.79	83.69	-3.43	3.91
PBE	96.79	89.65	87.24	84.08	-3.98	3.98
M06-L	96.32	88.96	84.49	81.75	-5.54	5.54
BP86	94.62	87.19	85.49	82.10	-6.07	6.07
mPWPW	94.58	87.22	85.26	81.87	-6.19	6.19
VSXC	90.11	87.22	81.16	83.72	-7.87	7.87
τ -HCTH	93.43	84.43	84.69	79.51	-7.91	7.91
HCTH	92.70	84.04	82.95	78.07	-8.98	8.98
TPSS	90.48	83.74	82.36	79.54	-9.39	9.39
OLYP	92.09	83.35	81.43	76.34	-10.12	10.12
BLYP ^a	90.31	82.64	81.09	77.50	-10.53	10.53
G96LYP	89.01	80.68	79.64	75.40	-12.24	12.24
	Nonlocal					
M05	94.47	86.99	86.32	82.77	-5.79	5.79
B3LYP ^a	91.58	85.01	82.58	80.06	-8.62	8.62
TPSSh	90.47	84.12	82.08	79.62	-9.35	9.35
Average					-7.73	7.76

^aData for these methods are taken from a paper by Izgorodina *et al.* (Ref. 75).

forms strikingly well (on a relative basis as compared to other local functionals and to B3LYP and TPSSh) for isomerization energies of conjugated systems, but less accurate for proton affinities of π conjugated systems. The former prob-

TABLE IX. Results for π systems. [The MG3S basis set and MP2/6-31+G(d,p) geometries are employed.]

Method	π IE3		PA-P5		PA-SB5		MMUE
	MSE	MUE	MSE	MUE	MSE	MUE	
	Local						
PBE	8.8	8.8	4.2	4.2	4.5	4.8	5.9
BB95	8.7	8.7	4.4	4.4	4.5	4.9	6.0
BP86	8.8	8.8	4.8	4.8	4.7	4.9	6.2
BLYP	8.7	8.7	4.9	5.5	4.3	4.8	6.3
M06-L	5.4	5.4	7.8	7.8	6.1	6.1	6.4
mPWPW	8.8	8.8	5.5	5.5	5.7	5.7	6.6
G96LYP	8.7	8.7	6.5	6.5	6.1	6.1	7.1
TPSS	8.4	8.4	8.4	8.4	7.4	7.4	8.1
HCTH	8.4	8.4	7.8	7.8	8.3	8.3	8.2
OLYP	8.7	8.7	8.3	8.3	7.7	7.7	8.2
VSXC	8.4	8.4	8.3	8.3	8.1	8.1	8.3
τ -HCTH	8.6	8.6	8.8	8.8	8.6	8.6	8.7
	Nonlocal						
M05	1.8	1.8	7.9	7.9	5.5	5.5	5.1
B3LYP	6.2	6.2	5.8	5.8	5.9	5.9	6.0
TPSSh	7.2	7.2	8.6	8.6	7.8	7.8	7.9
Average	7.7	7.7	6.8	6.8	6.3	6.4	7.0

lem provides a difficult test for most DFT methods,^{69,77} and the latter problem is related to the notorious problem of overestimation of polarizabilities and hyperpolarizabilities of conjugated systems by DFT.^{148,149} It seems very hard to treat the latter problem with local functionals.

V.G. Metal atom excitation energy

Table X presents the performance for the MAEE5 database of atomic excitation energies. Among the tested DFT methods, VSXC, M06-L, and OLYP give the best performance, whereas BB95 gives poor results for this database. In fact, the considerable improvement for this data set of M06-L over that obtained with M05 is a significant point in convincing us that the functional form has been improved, especially since no excitation energies were used in training.

V.H. Overall performance for energetic data

Table XI is a summary of the performance of the tested methods for all quantities studied in this paper. To construct Table XI, we first ported the final columns of Tables III-X. Then we computed the combined MUE (CMUE) column as a straight average of these eight mean unsigned errors. The average of all these averages is 5.72 kcal/mol, but the average of the eight columns range from 1.5 to 9.1 kcal/mol. Thus some columns of Table XI have a larger effect on CMUE than others. This is unavoidable and there can be no universally satisfactory way to combine errors for different kinds of properties. Nevertheless, we do tabulate one more figure of merit, called the weighted MUE (WMUE). For this column we scaled each of the first eight numerical columns of Table XI by (5.72 kcal/mol)/Average, where Average is the average for that column (as given in the last row). Thus, for example, every value in the main-group thermochemistry column is scaled by 5.72/1.85. The WMUE is the average of the scaled values. Encouragingly, the conclusions drawn from the CMUE and WMUE columns are similar. (One could try other, more sophisticated weightings, but our goal is to understand the trends, not to develop arbitrary statistics.)

The M06-L weighted mean unsigned error is better than that of all of the hybrid functionals tested, whereas none of the other local functional is better than any of them.

V.I. Bond lengths and frequencies

Table XII is a summary of the performance of the M06-L functional, as well as that of B3LYP, TPSS, BLYP, PBE, and VSXC, for the prediction of the bond lengths of 13 main-group molecules in the MGBL19 database and 13 metal-ligand bonds in the MBLB13/05 database. For the main-group database, Table XII shows that M06-L and VSXC give the smallest MUE for the prediction of bond lengths of main-group compounds; both outperform the B3LYP functional, and BLYP, PBE, and TPSS all give considerably worse geometries for the main-group database. In contrast, VSXC performs worst for the metal-ligand database, and the best performing functionals for this database are PBE, TPSS, and B3LYP. We also tabulated a quantity

TABLE X. Metal atom excitation energy (kcal/mol). [All calculated excitation energies are for the lowest excited state and include the relativistic contribution. The aug-cc-pVQZ basis sets are employed for Be, Mg, and Cu. The basis set for Pd is taken from Quintal (Ref. 94). The QZVP (Ref. 91) basis set is employed for Mn.]

	Be	Mg	Mn	Cu ⁺	Pd	MSE	MUE
Expt.	62.84	62.47	48.76	62.70	18.77		
Scalar rel. ^a	^b	^b	3.92 ^c	-10.38 ^d	^e		
S-O ^f	-0.01	-0.12	-0.66	-2.06	-2.26		
Total rel. ^g	-0.01	-0.12	3.26	-12.44	-2.26		
	Local						
VSXC	58.76	60.14	44.76	66.19	18.82	-1.4	2.8
M06-L	53.39	60.87	46.50	68.20	21.68	-3.3	4.0
OLYP	59.45	62.43	29.38	67.78	21.88	-2.9	6.2
HCTH	62.27	69.96	44.53	77.40	29.57	5.6	7.6
BLYP	56.81	65.46	25.99	56.52	15.77	-7.0	8.2
PBE	53.09	60.09	23.05	58.68	16.62	-8.8	8.8
BP86	54.83	60.81	21.37	58.31	15.75	-8.9	8.9
G96LYP	54.25	61.28	22.04	57.63	15.45	-9.0	9.0
TPSS	55.70	58.95	21.37	56.59	14.02	-9.8	9.8
mPWPW91	52.29	58.57	20.41	58.64	16.19	-9.9	9.9
τ -HCTH	64.94	72.14	51.44	89.37	34.44	11.4	11.4
BB95	54.78	60.52	22.66	41.22	19.31	-11.4	11.6
	Nonlocal						
B3LYP	56.61	64.11	35.24	52.37	15.20	-6.4	7.1
M05	64.30	67.41	57.14	71.21	34.11	7.7	7.7
TPSSh	55.78	58.69	26.55	54.40	14.02	-9.2	9.2
Average						-4.2	8.1

^aScalar relativistic effect.

^bAssumed to be negligible.

^cTaken from Raghavachari and Trucks (Ref. 97).

^dTaken from Martin and Hay (Ref. 64).

^eThe effect is taken into account by using a relativistic effective core potential.

^fSpin-orbit effect.

^gThis row, which is the sum of the previous two, is the total relativistic effect that has been added to all DFT results to produce the values in the 15 DFT rows of this table.

called AMUE in Table XII; this is an average of the errors over the two bond-length databases. Encouragingly, M06-L gives the smallest AMUE.

Table XIII presents the results for the prediction of harmonic frequencies. Overall the best performer is VSXC, followed by B3LYP and M06-L. M06-L has a mean unsigned error of 45 cm⁻¹ for frequencies greater than 2000 cm⁻¹ and a mean unsigned error of 33 cm⁻¹ for those under 2000 cm⁻¹.

We note that no geometries or frequencies were used in the parametrization of M06-L.

V.J. Comparison to a fifth-rung semiempirical functional

Recently, Grimme¹⁵⁰ proposed a fifth-rung semiempirical functional, called B2-PLYP, which employs the Kohn-Sham orbitals and eigenvalues as input for a standard MP2 perturbation correction (earlier fifth-rung functionals, that is, functionals that depend on the virtual orbitals, are discussed elsewhere^{67,117,151-153}). The two new semiempirical parameters in B2-PLYP were optimized against the G2/97 (Ref. 154) neutral test set. (Note that, also discussed by Perdew *et al.*,²⁰ there are five other empirical parameters in the B2-PLYP functional: one is in the Becke88 (Ref. 7) exchange

and four in the LYP correlation.⁸) Although B2-PLYP has less parameters than M06-L, it involves Hartree-Fock (HF) exchange, and it scales as N^5 (where N is the size of a system). We were asked to examine the performance of B2-PLYP in the context of the present work. We are happy to do so, but rather than add B2-PLYP to all tables, we will examine it for two representative databases¹⁵⁵ of main-group thermochemistry and thermochemical kinetics, for the prototype problem of π - π stacking in benzene dimer, and for a problem recently singled out¹⁵⁶ by Grimme. The representative databases are AE6 for atomization energies and BH6 for barrier heights. The data in each of these databases are subsets of the data in MGAE109 and BH76; in particular, they have been shown to be subsets for which performance is indicative of performance on the larger data bases.^{80,155}

The problem recently studied by Grimme and suggested by him to be “a mandatory benchmark for new density functionals” is the isomerization energy of tetramethylbutane to *n*-octane. Grimme found that the accurate result is +1.9 kcal/mol; three *ab initio* wave function calculations gave -11.5, +1.4, and +4.6 kcal/mol, B2-PLYP gave -3.5 kcal/mol, and seven density functionals gave values in the range from -9.9 to -5.5 kcal/mol. We find here that M06-L gives +0.6 kcal/mol. Grimme concluded that the

TABLE XI. Overall results.

Method	Thermochemistry TMUE	Barriers MMUE	Noncovalent MMMUE	Biological MMUE	TM MMUE	BDE MUE	π -system MMUE	Excitation MUE	CMUE ^a	WMUE ^b
Local										
M06-L	1.39	3.88	0.55	0.77	5.71	5.54	6.42	3.95	3.53	3.34
HCTH	1.89	4.81	1.26	4.11	9.19	8.98	8.17	7.56	5.75	5.63
PBE	3.09	8.64	1.14	2.44	10.23	3.98	5.93	8.80	5.53	5.64
TPSS	1.52	8.54	1.22	3.19	8.07	9.39	8.08	9.78	6.22	5.86
mPWPW	2.33	8.27	1.30	3.45	9.12	6.19	6.64	9.89	5.90	5.87
VSXC	1.17	5.00	2.39	7.19	8.56	7.87	8.30	2.79	5.41	5.93
τ -HCTH	1.54	6.15	1.20	3.73	13.78	7.91	8.65	11.36	6.79	6.20
BB95	2.50	7.94	1.67	3.98	11.65	3.91	6.00	11.63	6.16	6.24
BLYP	2.02	8.52	1.63	4.60	8.50	10.53	6.34	8.19	6.29	6.33
BP86	3.96	8.86	1.51	3.79	10.03	6.07	6.16	8.89	6.16	6.62
OLYP	1.43	5.53	2.62	7.12	7.69	10.12	8.24	6.20	6.12	6.62
G96LYP	2.40	7.03	2.75	7.05	7.61	12.24	7.09	8.98	6.89	7.51
Nonlocal										
M05	1.12	2.19	0.52	1.83	6.75	5.79	5.10	7.06	3.79	3.44
B3LYP	1.47	4.50	1.13	3.54	11.57	8.62	5.98	7.73	5.57	5.22
TPSSh	1.53	6.62	1.05	3.01	8.73	9.35	7.87	9.22	5.92	5.52
Average	1.85	6.43	1.46	3.99	9.15	7.77	7.00	8.13	5.72	5.72

^aCMUE denotes composite MUE, and it is the average of the eight previous columns.

^bWMUE denotes weighted MUE; see Sec. V.H for the definition of WMUE.

density functional results in his paper are “bad news” for standard Kohn-Sham DFT as it seems difficult to obtain the necessary information about electron pairs and their interactions from simple considerations of the electron density alone. Of course, we are not questioning the exactness and usefulness of the Hohenberg-Kohn theorems or the Kohn-Sham approach in general, but just the human ability to find accurate density functionals in practice. The better results with the new virtual orbital-dependent B2-PLYP functional that includes in part the necessary terms show, however, the way to go in the future.” We are not questioning the usefulness of fifth-rung functionals when expensive calculations are affordable, but it is encouraging that the less expensive M06-L calculations outperform B2-PLYP for this important benchmark. We note that Grimme’s paper was published after our paper was submitted.

Table XIV compares errors for the thermochemical AE6 (Ref. 155) database of six atomization energies, for the kinetics BH6 (Ref. 155) database of six barrier heights, for the binding energy¹⁵⁷ of parallel-displaced benzene dimer, and for the isomerization energy^{156,158} of octane. From Table IV, one can see that B2-PLYP significantly outperforms M06-L for barrier heights, which is not surprising because B2-PLYP contains 47% HF exchange, whereas M06-L is a local functional without any HF exchange. Furthermore, the comparisons already presented show that barrier heights are the one area where M06-L is significantly outperformed by new-generation density functionals such as M05 (Refs. 17 and 44) and M05-2X (Ref. 17) (although not by popular functionals such as B3LYP). Turning to the other tests in Table XIV, we see that the expensive B2-PLYP outperforms M06-L by 11% for thermochemistry, as shown by the MUE for the AE6 database, but it is not as accurate for the binding energy of

benzene dimer and the isomerization energy of octane in Table XIV, which are both notorious examples for which many density functionals fail.^{70,156}

Furthermore, Table XIV shows that B2-PLYP is one order of magnitude more expensive than M06-L, as measured by the computer time for a benzene dimer calculation, and as one increases system size, this ratio of computer times will become more dramatic because M06-L scales as N^3 , whereas B2-PLYP scales as N^5 . Note that the calculations in Ref. 156 employed a polarized quadruple zeta basis for B2-PLYP, which requires larger basis sets than M06-L; nevertheless, the timing in Table XIV is for the MG3S polarized triple zeta basis set (using a polarized quadruple zeta basis set would make the timing comparison less favorable for B2-PLYP).

VI. CONCLUDING REMARKS

This paper presents a new local meta-GGA exchange-correlation functional, M06-L, for thermochemistry, thermochemical kinetics, and noncovalent interactions. This functional builds on earlier work of several authors,^{6,11,12,15,17,44,49,100–103} and it is designed to incorporate spin kinetic energy density in a balanced way in the exchange and correlation functionals. In addition, it satisfies the uniform-electron-gas limit, and it is self-correlation-free. The M06-L functional was comparatively assessed against 22 diverse energetic databases, a bond-length test set, and a set of 36 harmonic vibrational frequencies.

The assessments include main-group chemistry, transition metal chemistry, and biological chemistry. From these assessments and from comparison with results for 14 functionals in the literature, we draw the conclusion, based on an analysis of mean unsigned errors, that the M06-L functional

TABLE XII. Performance for the prediction of bond lengths (Å) in the MGBL19 and MLBL13/05 databases. (The MG35 basis set is employed for all functionals in this table.)

Bond lengths	Expt. ^a	M06-L	VSXC	B3LYP	TPSS	PBE	BLYP	
MGBL19 database								
H ₂	0.741	0.743	0.741	0.743	0.743	0.751	0.747	
CH ₄	1.086	1.086	1.087	1.088	1.092	1.096	1.094	
NH ₃	1.012	1.012	1.011	1.013	1.019	1.021	1.021	
H ₂ O	0.957	0.958	0.958	0.961	0.968	0.969	0.971	
HF	0.917	0.917	0.919	0.922	0.929	0.930	0.933	
CO	1.128	1.128	1.133	1.125	1.135	1.137	1.137	
N ₂	1.098	1.096	1.098	1.091	1.100	1.103	1.103	
F ₂	1.412	1.405	1.414	1.396	1.415	1.412	1.432	
C ₂ H ₂	C–C	1.203	1.198	1.203	1.196	1.202	1.207	1.206
	C–H	1.063	1.062	1.063	1.062	1.065	1.071	1.067
HCN	C–H	1.065	1.067	1.067	1.066	1.070	1.075	1.072
	C–N	1.153	1.150	1.154	1.146	1.155	1.158	1.157
H ₂ CO	C–H	1.102	1.111	1.109	1.106	1.110	1.117	1.114
	C–O	1.203	1.195	1.204	1.199	1.208	1.209	1.212
CO ₂	1.160	1.159	1.165	1.160	1.169	1.171	1.173	
N ₂ O	N–N	1.128	1.129	1.131	1.121	1.134	1.138	1.139
	N–O	1.184	1.178	1.184	1.184	1.191	1.190	1.198
OH	0.970	0.971	0.972	0.974	0.982	0.983	0.985	
Cl ₂	1.988	1.980	2.012	2.016	2.016	2.011	2.046	
MSE ^b		−0.0014	0.0029	−0.0024	0.0070	0.0094	0.0125	
MUE ^c		0.0030	0.0031	0.0055	0.0071	0.0094	0.0125	
MLBL13/05 database								
AgH	1.618	1.644	1.630	1.637	1.634	1.627	1.634	
BeO	1.331	1.321	1.345	1.321	1.336	1.340	1.340	
CoH	1.531	1.538	1.528	1.532	1.530	1.526	1.529	
CoO ⁺	1.646	1.637	1.643	1.636	1.631	1.627	1.644	
FeH	1.610	1.584	1.579	1.573	1.571	1.570	1.572	
FeO	1.616	1.611	1.628	1.612	1.606	1.606	1.621	
FeS	2.013	1.998	2.021	2.028	2.011	2.011	2.030	
LiCl	2.021	2.025	2.042	2.024	2.030	2.028	2.034	
LiO	1.689	1.695	1.719	1.690	1.701	1.704	1.705	
MgO	1.749	1.733	1.751	1.738	1.745	1.749	1.755	
RhC	1.613	1.616	1.622	1.608	1.623	1.621	1.633	
VO	1.589	1.583	1.594	1.580	1.590	1.585	1.598	
VS	2.048	2.056	2.072	2.058	2.054	2.050	2.066	
MSE ^b		−0.003	0.008	−0.003	−0.001	−0.002	0.007	
MUE ^c		0.011	0.013	0.010	0.010	0.010	0.013	
Both databases								
AMUE ^d		0.007	0.008	0.008	0.009	0.010	0.013	

^aFrom Ref. 48.^bMSE denotes mean signed error.^cMUE denotes mean unsigned error.^dAMUE is the average of the two MUEs in this table.

gives the best overall performance of any functional for a combination of thermochemistry, thermochemical kinetics, metallochemical and noncovalent interactions, bond lengths, and vibrational frequencies. It is striking that it outperforms the second best local functional by almost a factor of 2, even though the uniform-electron-gas limit was sacrificed in that functional to improve empirical performance. Furthermore, Tables III–XIII show that M06-L does well by being the best or reasonably close to the best local functional in in almost all categories, the main (and troubling) exceptions being pro-

ton affinities of conjugated π bonded systems and electron affinities. We especially recommend the M06-L functional for the study of systems involving transition metal bonding, since the performance is better than that of all other functionals except M05 by a large margin. (It outperforms M05 by only 19%.)

The new M06-L functional not only excels in comparison to other local functionals, its average overall performance for energetic quantities (as measured by overall performance on both of the last two columns of Table XI) is

TABLE XIII. Performance for the prediction of harmonic frequencies (cm^{-1}). (The MG3S basis set is employed for all functionals in this table. Although all values are rounded to the nearest integer in this table, mean errors were computed from unrounded data.)

Molecule	Mode	Best est. ^a	VSXC	B3LYP	M06-L	TPSS	PBE	BLYP
H ₂	ω_1	4401	4412	4421	4324	4419	4319	4349
CH ₄	ω_1	1367	1320	1345	1334	1343	1289	1313
	ω_2	1583	1550	1562	1564	1561	1512	1527
	ω_3	3026	3034	3034	3042	3003	2976	2967
	ω_4	3157	3156	3135	3180	3105	3089	3061
NH ₃	ω_1	3478	3489	3483	3492	3406	3401	3374
	ω_2	1084	1025	1017	1073	1054	1010	1011
	ω_3	3597	3618	3603	3627	3523	3524	3489
	ω_4	1684	1671	1669	1688	1667	1624	1634
H ₂ O	ω_1	1648	1642	1631	1665	1632	1597	1599
	ω_2	3832	3827	3819	3860	3719	3719	3676
	ω_3	3942	3937	3922	3976	3824	3825	3779
HF	ω_1	4138	4102	4096	4169	3993	3987	3938
CO	ω_1	2170	2165	2214	2204	2140	2130	2116
N ₂	ω_1	2359	2394	2445	2407	2365	2347	2329
F ₂	ω_1	917	974	1045	979	1001	992	955
C ₂ H ₂	ω_1	624	616	665	660	615	607	603
	ω_2	747	760	765	797	763	733	733
	ω_3	2008	2039	2067	2062	2028	2009	2003
	ω_4	3415	3420	3409	3401	3380	3342	3339
	ω_5	3495	3518	3509	3504	3478	3440	3435
HCN	ω_1	727	753	764	789	747	729	724
	ω_2	2127	2154	2199	2175	2136	2119	2110
	ω_3	3443	3437	3439	3415	3400	3360	3356
H ₂ CO	ω_1	2937	2857	2888	2847	2851	2798	2791
	ω_2	1778	1794	1817	1853	1768	1764	1738
	ω_3	1544	1518	1531	1524	1518	1479	1490
	ω_4	1188	1185	1200	1196	1178	1150	1154
	ω_5	3012	2906	2943	2895	2899	2842	2834
	ω_6	1269	1246	1264	1257	1245	1219	1226
CO ₂	ω_1	673	668	679	690	652	647	640
	ω_2	1354	1345	1371	1385	1325	1322	1306
	ω_3	2397	2401	2408	2473	2358	2359	2319
N ₂ O	ω_1	596	617	620	638	599	597	579
	ω_2	1298	1322	1326	1356	1290	1300	1257
	ω_3	2282	2320	2340	2367	2269	2279	2219
All	MSE ^a	MSE	-3	10	16	-29	-52	-64
	MUE ^c	MUE	24	31	39	39	56	67
>2000 cm^{-1}	MSE ^b	MSE	-1	8	11	-48	-71	-91
	MUE ^c	MUE	24	32	45	54	71	91
<2000 cm^{-1}	MSE ^b	MSE	-4	11	22	-7	-30	-35
	MUE ^c	MUE	23	30	33	22	40	39

^aFrom Ref. 81.

^bMSE denotes mean signed error.

^cMUE denotes mean unsigned error.

better than any of the three hybrid functionals we tested, including the best previous broadly accurate functional (M05) and the very popular B3LYP one. Since the B3LYP functional has become a generally accepted standard for a broadly accurate functional, it is useful to mention that the mean unsigned error for the local M06-L functional is

smaller than that for the hybrid B3LYP functional by a factor of 1.5 if measured by either of the last two columns of Table XI.

From a fundamental point of view, as pointed by Becke¹⁰³ in a related context, it is important to carry out studies like the present one to learn the limit of accuracy of

TABLE XIV. Comparison to a fifth-rung functional.

	M06-L	B2-PLYP
MUE of AE6 ^a	3.8 ^b	3.4 ^b
MUE of BH6 ^a	4.3 ^b	1.9 ^b
UE of D_e of benzene dimer ^{c,d}	0.4	2.5
UE of the isomerization energy of octane ^{c,e}	1.3	5.4 ^f
Cost ^g	324	3906

^aMUE is mean unsigned error in kcal/mol.

^bThe MG3S basis set and QCISD/MG3 geometries were employed for calculations on the AE6 and BH6 databases. To convert the AE6 results in this table to a per bond basis (which was not done), one would divide by 4.83.

^cUE is unsigned error in kcal/mol.

^dThe best estimate D_e for the parallel displaced configuration is 2.8 kcal/mol (Ref. 157); D_e for M06-L is 2.4 kcal/mol. The UE for the B2-PLYP method is taken from Ref. 50.

^eThe best estimate is 1.9 kcal/mol for the zero-point exclusive energy of *n*-octane minus that of tetramethylbutane (Refs. 156 and 158); M06-L gives 0.6 Kcal/mol, whereas B2-PLYP predicts a value of -3.5 kcal/mol (Ref. 156). The basis set used for the octane isomerization calculation is MG3S for M06-L and the level is cQZV3P//MP2/TZV(*d,p*) for B2-PLYP.

^fTaken from Ref. 156.

^gCost is measured as the computer time (in arbitrary units) for a single-point energy calculation of the benzene dimer (sandwich) with the 6-311+G(2df,2p) basis set on a single processor of an Altix 3700 computer.

particular kinds of density functionals, in his case hybrid GGAs, and in our case meta-GGAs without Hartree-Fock exchange. As compared to the earlier tests of Becke and many others that focused exclusively on main-group thermochemistry of small molecules, this study includes barriers, noncovalent interactions, transition metals, and larger systems. Our conclusions overturn the widespread belief that local functionals are not capable of improving on the broad average accuracy of the best current hybrid functionals.

It is very encouraging that we succeeded in developing a local density functional with very broad applicability. From a practical point of view, the new high-performance local functional presented here, because of the computational efficiencies possible with local functionals, should be immediately useful for a myriad of applications involving complex systems, such as organometallic and inorganometallic chemistry, biological problems, soft materials, and molecular solids (including ice). We especially stress its usefulness for condensed-phase systems and molecular recognition problems (including supramolecular chemistry and protein assemblies) when noncovalent interactions are very important and hybrid functionals are prohibitively expensive.

ACKNOWLEDGMENTS

The authors are grateful to Nathan Schultz for helpful discussions. This work was supported in part by the U. S. Department of Energy, Office of Basic Energy Sciences under Grant No. DE-FG02-86ER13579 (kinetics), by the National Science Foundation under Grant Nos. CHE03-49122 (complex systems) and ITR04-28774 (planetary ice), and by the Office of Naval Research under Award No. N00014-05-0538 (tools and software).

¹W. Kohn and L. J. Sham, Phys. Rev. **140**, 1133 (1965).

²S. H. Vosko, L. Wilk, and M. Nusair, Can. J. Phys. **58**, 1200 (1980).

³J. P. Perdew and Y. Wang, Phys. Rev. B **45**, 13244 (1992).

⁴D. C. Langreth and M. J. Mehl, Phys. Rev. B **28**, 1809 (1983).

⁵J. P. Perdew, Phys. Rev. B **33**, 8822 (1986).

⁶A. D. Becke, J. Chem. Phys. **84**, 4524 (1986).

⁷A. D. Becke, Phys. Rev. A **38**, 3098 (1988).

⁸C. Lee, W. Yang, and R. G. Parr, Phys. Rev. B **37**, 785 (1988).

⁹T. Ziegler, Chem. Rev. (Washington, D.C.) **91**, 651 (1991).

¹⁰M. Filatov and W. Thiel, Mol. Phys. **91**, 847 (1997).

¹¹A. D. Becke, J. Chem. Phys. **104**, 1040 (1996).

¹²T. V. Voorhis and G. E. Scuseria, J. Chem. Phys. **109**, 400 (1998).

¹³J. Tao, J. P. Perdew, V. N. Staroverov, and G. E. Scuseria, Phys. Rev. Lett. **91**, 146401 (2003).

¹⁴A. D. Becke, J. Chem. Phys. **98**, 5648 (1993).

¹⁵J. P. Perdew, M. Ernzerhof, and K. Burke, J. Chem. Phys. **105**, 9982 (1996).

¹⁶C. Adamo and V. Barone, J. Chem. Phys. **108**, 664 (1998).

¹⁷Y. Zhao, N. E. Schultz, and D. G. Truhlar, J. Chem. Theory Comput. **2**, 364 (2006).

¹⁸P. Hohenberg and W. Kohn, Phys. Rev. **136**, 864 (1964).

¹⁹J. P. Perdew and A. Zunger, Phys. Rev. B **23**, 5048 (1981).

²⁰J. P. Perdew, A. Ruzsinszky, J. Tao, V. N. Staroverov, G. E. Scuseria, and G. I. Csonka, J. Chem. Phys. **123**, 062201 (2005).

²¹E. J. Baerends, D. E. Ellis, and P. Ros, Chem. Phys. **2**, 41 (1973).

²²B. I. Dunlap, J. W. D. Connolly, and J. R. Sabin, J. Chem. Phys. **71**, 3396 (1979).

²³O. Vahtras, J. Almlöf, and M. W. Feyereisen, Chem. Phys. Lett. **213**, 514 (1993).

²⁴R. A. Kendall, E. Apra, D. E. Bernholdt *et al.*, Comput. Phys. Commun. **128**, 260 (2000).

²⁵G. Te Velde, F. M. Bickelhaupt, E. J. Baerends, C. Fonseca Guerra, S. J. A. Van Gisbergen, J. G. Snijders, and T. Ziegler, J. Comput. Chem. **22**, 931 (2001).

²⁶J. VandeVondele, M. Krack, F. Mohamed, M. Parrinello, T. Chassaing, and J. Hutter, Comput. Phys. Commun. **167**, 103 (2005).

²⁷Y. Jung, A. Sodt, P. M. W. Gill, and M. Head-Gordon, Proc. Natl. Acad. Sci. U.S.A. **102**, 6692 (2005).

²⁸K. Eichkorn, O. Treutler, H. Oehm, M. Haeser, and R. Ahlrichs, Chem. Phys. Lett. **240**, 283 (1995).

²⁹K. Eichkorn, F. Weigend, O. Treutler, and R. Ahlrichs, Theor. Chem. Acc. **97**, 119 (1997).

³⁰L. Füstí-Molnár and P. Pulay, THEOCHEM **666-667**, 25 (2003).

³¹Y.-B. Wang and Z. Lin, J. Am. Chem. Soc. **25**, 6072 (2003).

³²M. J. Frisch, G. W. Trucks, H. B. Schlegel *et al.*, GAUSSIAN03, Gaussian, Inc., Pittsburgh, PA, 2003.

³³G. E. Scuseria and V. N. Staroverov, in *Theory and Application of Computational Chemistry: The First 40 Years*, edited by C. E. Dykstra, G. Frenking, K. S. Kim, and G. E. Scuseria (Elsevier, Amsterdam, 2005), p. 669.

³⁴J. Heyd and G. E. Scuseria, J. Chem. Phys. **120**, 7274 (2004).

³⁵J. Heyd and G. E. Scuseria, J. Chem. Phys. **121**, 1187 (2004).

³⁶J. Heyd, J. E. Peralta, G. E. Scuseria, and R. L. Martin, J. Chem. Phys. **123**, 174101 (2005).

³⁷J. Paier, M. Marsman, K. Hummer, G. Kresse, I. C. Gerber, and J. G. Ángyán, J. Chem. Phys. **124**, 154709 (2006).

³⁸N. Schultz, Y. Zhao, and D. G. Truhlar, J. Phys. Chem. A **109**, 4388 (2005).

³⁹N. Schultz, Y. Zhao, and D. G. Truhlar, J. Phys. Chem. A **109**, 11127 (2005).

⁴⁰O. V. Gritsenko, P. R. T. Schipper, and E. J. Baerends, J. Chem. Phys. **107**, 5007 (1997).

⁴¹N. C. Handy and A. J. Cohen, Mol. Phys. **99**, 403 (2001).

⁴²A. D. Becke, J. Chem. Phys. **119**, 2972 (2003).

⁴³A. D. Becke, J. Chem. Phys. **122**, 064101 (2005).

⁴⁴Y. Zhao, N. E. Schultz, and D. G. Truhlar, J. Chem. Phys. **123**, 161103 (2005). Note that in this communication we interchanged $c_{C\alpha\beta,i}$ and $c_{C\sigma\sigma,i}$ in Table I. In addition, "reduced density x_σ " before Eq. (1) should read "reduced density gradient x_σ ".

⁴⁵P. M. W. Gill, Mol. Phys. **89**, 433 (1996).

⁴⁶J. P. Perdew, K. Burke, and M. Ernzerhof, Phys. Rev. Lett. **77**, 3865 (1996).

⁴⁷J. P. Perdew, in *Electronic Structure of Solids '91*, edited by P. Ziesche and H. Eschig (Akademie, Berlin, 1991), p. 11.

⁴⁸F. A. Hamprecht, A. J. Cohen, D. J. Tozer, and N. C. Handy, J. Chem. Phys. **109**, 6264 (1998).

⁴⁹A. D. Boese and N. C. Handy, J. Chem. Phys. **116**, 9559 (2002).

- ⁵⁰ P. J. Stephens, F. J. Devlin, C. F. Chabalowski, and M. J. Frisch, *J. Phys. Chem.* **98**, 11623 (1994).
- ⁵¹ V. N. Staroverov, G. E. Scuseria, J. Tao, and J. P. Perdew, *J. Chem. Phys.* **119**, 12129 (2003).
- ⁵² T. Lovell, R. Stranger, and J. E. McGrady, *Inorg. Chem.* **40**, 39 (2001).
- ⁵³ M. Reiher, O. Salomon, and B. A. Hess, *Theor. Chem. Acc.* **107**, 48 (2001).
- ⁵⁴ J. N. Harvey, R. Poli, and K. M. Smith, *Coord. Chem. Rev.* **238–239**, 347 (2003).
- ⁵⁵ J.-L. Carreón-Macedo and J. N. Harvey, *J. Am. Chem. Soc.* **126**, 5789 (2004).
- ⁵⁶ R. Poli, *J. Organomet. Chem.* **689**, 4291 (2004).
- ⁵⁷ C. Hermann, J. Nevgebauer, J. A. Gladysz, and M. Reiher, *Inorg. Chem.* **44**, 6175 (2005).
- ⁵⁸ D. M. Jenkins and J. C. Peters, *J. Am. Chem. Soc.* **127**, 7148 (2005).
- ⁵⁹ F. Neese, *J. Inorg. Biochem.* **100**, 716 (2006).
- ⁶⁰ J.-L. Carreón-Macedo and J. N. Harvey, *Phys. Chem. Chem. Phys.* **8**, 93 (2006).
- ⁶¹ M. P. Mehn, S. D. Brown, T. K. Paine, W. W. Brennessel, C. J. Cramer, J. C. Peters, and L. J. Que, *Dalton Trans.* **2006**, 1347 (2006).
- ⁶² I. Rudra, Q. Wu, and T. V. Voorhis, *J. Chem. Phys.* **124**, 024103 (2006).
- ⁶³ M. C. Holthausen, *J. Comput. Chem.* **26**, 1505 (2005).
- ⁶⁴ R. L. Martin and P. J. Hay, *J. Chem. Phys.* **75**, 4539 (1981).
- ⁶⁵ B. J. Lynch, Y. Zhao, and D. G. Truhlar, *J. Phys. Chem. A* **107**, 1384 (2003).
- ⁶⁶ B. J. Lynch and D. G. Truhlar, *J. Phys. Chem. A* **107**, 3898 (2003).
- ⁶⁷ Y. Zhao, B. J. Lynch, and D. G. Truhlar, *Phys. Chem. Chem. Phys.* **7**, 43 (2005).
- ⁶⁸ Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A* **109**, 5656 (2005).
- ⁶⁹ Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A* **110**, 10478 (2006).
- ⁷⁰ Y. Zhao, N. González-García, and D. G. Truhlar, *J. Phys. Chem. A* **109**, 2012 (2005).
- ⁷¹ Y. Zhao and D. G. Truhlar, *J. Chem. Theory Comput.* **1**, 415 (2005).
- ⁷² P. Jurecka, J. Sponer, J. Cerny, and P. Hobza, *Phys. Chem. Chem. Phys.* **8**, 1985 (2006).
- ⁷³ Y. Zhao and D. G. Truhlar, *J. Chem. Phys.* **124**, 224105 (2006).
- ⁷⁴ F. Furche and J. P. Perdew, *J. Chem. Phys.* **124**, 044103 (2006).
- ⁷⁵ E. I. Izgorodina, M. L. Coote, and L. Radom, *J. Phys. Chem. A* **109**, 7558 (2005).
- ⁷⁶ A. P. Scott and L. Radom, *J. Phys. Chem.* **100**, 16502 (1996).
- ⁷⁷ H. L. Woodcock, H. F. Schaefer, and P. R. Schreiner, *J. Phys. Chem. A* **106**, 11923 (2002).
- ⁷⁸ See EPAPS Document No. E-JCPSA6-125-309642 for all previously published databases. This document can be reached via a direct link in the online article's HTML reference section or via the EPAPS homepage (<http://www.aip.org/pubservs/epaps.html>).
- ⁷⁹ P. L. Fast, J. Corchado, M. L. Sanchez, and D. G. Truhlar, *J. Phys. Chem. A* **103**, 3139 (1999).
- ⁸⁰ Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A* **108**, 6908 (2004).
- ⁸¹ J. M. L. Martin, J. El-Yazal, and J.-P. Francois, *Mol. Phys.* **86**, 1437 (1995).
- ⁸² J. M. L. Martin, *J. Chem. Phys.* **100**, 8186 (1994).
- ⁸³ J. A. Pople, M. Head-Gordon, and K. Raghavachari, *J. Chem. Phys.* **87**, 5968 (1987).
- ⁸⁴ P. L. Fast, M. L. Sanchez, and D. G. Truhlar, *Chem. Phys. Lett.* **306**, 407 (1999).
- ⁸⁵ L. A. Curtiss, P. C. Redfern, K. Raghavachari, V. Rassolov, and J. A. Pople, *J. Chem. Phys.* **110**, 4703 (1999).
- ⁸⁶ L. A. Curtiss, K. Raghavachari, P. C. Redfern, V. Rassolov, and J. A. Pople, *J. Chem. Phys.* **109**, 7764 (1998).
- ⁸⁷ M. J. Frisch, J. A. Pople, and J. S. Binkley, *J. Chem. Phys.* **80**, 3265 (1984).
- ⁸⁸ W. J. Hehre, L. Radom, P. v. R. Schleyer, and J. A. Pople, *Ab Initio Molecular Orbital Theory*, 1st ed. (Wiley, New York, 1986).
- ⁸⁹ P. L. Fast and D. G. Truhlar, *J. Phys. Chem. A* **104**, 6111 (2000).
- ⁹⁰ M. O. Sinnokrot and C. D. Sherrill, *J. Phys. Chem. A* **108**, 10200 (2004).
- ⁹¹ F. Weigend, F. Furche, and R. Ahlrichs, *J. Chem. Phys.* **119**, 12753 (2003).
- ⁹² D. E. Woon and J. T. H. Dunning, *J. Chem. Phys.* **98**, 1358 (1993).
- ⁹³ N. B. Balabanov and K. A. Peterson, *J. Chem. Phys.* **123**, 064107 (2005).
- ⁹⁴ M. M. Quintal, A. Karton, M. A. Iron, A. D. Boese, and J. M. L. Martin, *J. Phys. Chem. A* **110**, 709 (2006).
- ⁹⁵ W. J. Stevens, M. Krauss, H. Basch, and P. G. Jasien, *Can. J. Chem.* **70**, 612 (1992).
- ⁹⁶ D. Andrae, U. Haeussermann, M. Dolg, H. Stoll, and H. Preuss, *Theor. Chim. Acta* **77**, 123 (1990).
- ⁹⁷ K. Raghavachari and G. W. Trucks, *J. Chem. Phys.* **91**, 1062 (1989).
- ⁹⁸ S. F. Boys and F. Bernardi, *Mol. Phys.* **19**, 553 (1970).
- ⁹⁹ D. W. Schwenke and D. G. Truhlar, *J. Chem. Phys.* **82**, 2418 (1985); **86**, 3760(E) (1987).
- ¹⁰⁰ H. Stoll, C. M. E. Pavlidou, and H. Preuss, *Theor. Chim. Acta* **49**, 143 (1978).
- ¹⁰¹ A. D. Becke, *J. Chem. Phys.* **109**, 2092 (1998).
- ¹⁰² A. D. Becke, *J. Chem. Phys.* **112**, 4020 (2000).
- ¹⁰³ A. D. Becke, *J. Chem. Phys.* **107**, 8554 (1997).
- ¹⁰⁴ A. D. Rabuck and G. E. Scuseria, *Chem. Phys. Lett.* **309**, 450 (1999).
- ¹⁰⁵ Y. Zhao, J. Pu, B. J. Lynch, and D. G. Truhlar, *Phys. Chem. Chem. Phys.* **6**, 673 (2004).
- ¹⁰⁶ E. R. Johnson, R. A. Wolkow, and G. A. DiLabio, *Chem. Phys. Lett.* **394**, 334 (2004).
- ¹⁰⁷ J. C. Sancho-Garcia and J. Cornil, *J. Chem. Phys.* **121**, 3096 (2004).
- ¹⁰⁸ Y. Zhao and D. G. Truhlar, *J. Chem. Theory Comput.* (in press).
- ¹⁰⁹ G. C. L. Chan, D. J. Tozer, and N. C. Handy, *J. Chem. Phys.* **107**, 1 (1997).
- ¹¹⁰ T. V. Voorhis and G. E. Scuseria, *Mol. Phys.* **92**, 601 (1997).
- ¹¹¹ J. W. Negele and D. Vautherin, *Phys. Rev. C* **5**, 1472 (1972).
- ¹¹² S. J. Chakravorty, S. R. Gwaltney, E. R. Davidson, F. A. Parpia, and C. F. Fischer, *Phys. Rev. A* **47**, 3649 (1993).
- ¹¹³ Y. Zhao, B. J. Lynch, and D. G. Truhlar, *J. Phys. Chem. A* **108**, 2715 (2004).
- ¹¹⁴ P. J. Wilson, T. J. Bradley, and D. J. Tozer, *J. Chem. Phys.* **115**, 9233 (2001).
- ¹¹⁵ X. Xu and W. A. Goddard, *Proc. Natl. Acad. Sci. U.S.A.* **101**, 2673 (2004).
- ¹¹⁶ A. D. Boese and J. M. L. Martin, *J. Chem. Phys.* **121**, 3405 (2004).
- ¹¹⁷ Y. Zhao, B. J. Lynch, and D. G. Truhlar, *J. Phys. Chem. A* **108**, 4786 (2004).
- ¹¹⁸ I. W. Hamley, *Angew. Chem., Int. Ed.* **42**, 1692 (2003).
- ¹¹⁹ G. K. Ghosh, *Bull. Mater. Sci.* **26**, 3 (2003).
- ¹²⁰ T. Kato, N. Misoshita, and K. Kishimoto, *Angew. Chem., Int. Ed.* **45**, 38 (2005).
- ¹²¹ J. Wu, *AIChE J.* **52**, 1169 (2006).
- ¹²² M. Lehmann, C. Köhn, H. Meier, S. Renker, and A. Oehlhof, *J. Mater. Chem.* **16**, 441 (2006).
- ¹²³ R. Holyst, *Soft Matter* **1**, 329 (2005).
- ¹²⁴ A. Casey, G. F. Fassell, M. McNamara, H. J. Byrne, and G. Chambers, *Synth. Met.* **153**, 357 (2005).
- ¹²⁵ M. R. Wilson, *Int. Rev. Phys. Chem.* **24**, 421 (2005).
- ¹²⁶ E. Ruiz, D. R. Salahub, and A. Vela, *J. Am. Chem. Soc.* **117**, 1141 (1995).
- ¹²⁷ E. Ruiz, D. R. Salahub, and A. Vela, *J. Phys. Chem.* **100**, 12265 (1996).
- ¹²⁸ E. T. Kool, J. C. Morales, and K. M. Guckian, *Angew. Chem., Int. Ed.* **39**, 990 (2000).
- ¹²⁹ P. Barthelemy, S. J. Lee, and M. Grinstaff, *Pure Appl. Chem.* **77**, 2133 (2005).
- ¹³⁰ J. Vondrásek, L. Bendová, V. Klusák, and P. Hobza, *J. Am. Chem. Soc.* **127**, 2615 (2005).
- ¹³¹ H. Mansikkamaeki, M. Nissinen, and K. Rissanen, *Angew. Chem., Int. Ed.* **43**, 1243 (2004).
- ¹³² M. Vázquez, M. R. Bermejo, M. Licchelli, A. M. González-Noya, R. M. Pedrido, C. Sangregorio, L. Sorace, A. M. García-Deibe, and J. Santmartín, *Eur. J. Inorg. Chem.* **17**, 3479 (2005).
- ¹³³ J. Hautman and M. L. Klein, *NATO ASI Ser., Ser. E* **205**, 395 (1991).
- ¹³⁴ K. D. Karlin, *Science* **261**, 701 (1993).
- ¹³⁵ R. H. Crabtree, *The Organometallic Chemistry of the Transition Metals*, 2nd ed. (Wiley, New York, 1994).
- ¹³⁶ S. M. George, *Chem. Rev. (Washington, D.C.)* **95**, 475 (1995).
- ¹³⁷ G. A. Somorjai, *Chem. Rev. (Washington, D.C.)* **96**, 1223 (1995).
- ¹³⁸ M. A. Ratner, B. Davis, M. Kemp, V. Mujica, A. Roitberg, and S. Yaliraki, *Ann. N.Y. Acad. Sci.* **852**, 22 (1998).
- ¹³⁹ *Transition State Modeling for Catalysis*, edited by D. G. Truhlar and K. Morokuma (American Chemical Society, Washington, DC, 1999).
- ¹⁴⁰ E. R. Davidson, *Chem. Rev. (Washington, D.C.)* **100**, 351 (2000).
- ¹⁴¹ P. E. M. Siegbahn and M. R. A. Blomberg, *Chem. Rev. (Washington, D.C.)* **100**, 421 (2000).
- ¹⁴² J. A. Gladysz, *Chem. Rev. (Washington, D.C.)* **100**, 1167 (2000).
- ¹⁴³ *Handbook on Metalloproteins*, edited by I. Bertini, A. Sigel, and H. Sigel (Dekker, New York, 2001).

- ¹⁴⁴ A. K. Rappe, W. M. Skiff, and C. J. Casewit, Chem. Rev. (Washington, D.C.) **100**, 1435 (2000).
- ¹⁴⁵ C. Coperat, M. Chabonas, R. P. Saint-Arromon, and J.-M. Basset, Angew. Chem., Int. Ed. **42**, 156 (2003).
- ¹⁴⁶ G. Cavgliasso and R. Stranger, Inorg. Chem. **44**, 5081 (2005).
- ¹⁴⁷ L. A. Curtiss, K. Raghavachari, P. C. Redfern, and J. A. Pople, J. Chem. Phys. **112**, 7374 (2000).
- ¹⁴⁸ B. Champagne, E. A. Perpete, S. J. A. van Gisbergen, E.-J. Baerends, J. G. Snijders, C. Soubra-Ghaoui, K. A. Robins, and B. Kirtman, J. Chem. Phys. **109**, 10489 (1998).
- ¹⁴⁹ B. Champagne, E. A. Perpete, D. Jacquemin, S. J. A. V. Gisbergen, E.-J. Baerends, C. Soubra-Ghaoui, K. A. Robins, and B. Kirtman, J. Phys. Chem. A **104**, 4755 (2000).
- ¹⁵⁰ S. Grimme, J. Chem. Phys. **124**, 034108 (2006).
- ¹⁵¹ I. Grabowski, S. Hirata, S. Ivanov, and R. J. Bartlett, J. Chem. Phys. **116**, 4415 (2002).
- ¹⁵² R. Pollet, A. Savin, T. Leininger, and H. Stoll, J. Chem. Phys. **116**, 1250 (2002).
- ¹⁵³ P. Mori-Sanchez, Q. Wu, and W. Yang, J. Chem. Phys. **123**, 62204 (2005).
- ¹⁵⁴ L. A. Curtiss, K. Raghavachari, P. C. Redfern, and J. A. Pople, J. Chem. Phys. **106**, 1063 (1997).
- ¹⁵⁵ B. J. Lynch and D. G. Truhlar, J. Phys. Chem. A **107**, 8996 (2003) **108**, 1460 (E) (2004).
- ¹⁵⁶ S. Grimme, Angew. Chem., Int. Ed. **45**, 4460 (2006).
- ¹⁵⁷ M. O. Sinnokrot, E. F. Valeev, and C. D. Sherrill, J. Am. Chem. Soc. **124**, 10887 (2002).
- ¹⁵⁸ NIST Standard Reference Database. <http://webbook.nist.gov/chemistry/>.